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List number of hours worked on assignment must match State Requirement. _____

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I affirm that I personally completed the entire text of the course. I also affirm that I completed the exam without assistance from any outside source.

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Grading Information

In order to maintain the integrity of our courses we do not distribute test scores, percentages or questions missed.

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Modern Disinfection CEU Course Answer Key

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This course contains general EPA's SDWA federal rule requirements. Please be aware that each state implements water / sampling procedures/safety/ environmental / building regulations that may be more stringent than EPA's regulations. Check with your state environmental/health agency for more information. These rules change frequently and are often difficult to interpret and follow. Be careful to be in compliance with your agency and do not follow this course for proper compliance.

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Modern Disinfection CEU Course Assignment

The Assignment (Exam) is also available in Word on the Internet for your Convenience, please visit www.ABCTLC.com and download the assignment and e-mail it back to TLC.

You will have 90 days from the start of this course to complete in order to receive your Professional Development Hours (**PDHs**) or Continuing Education Unit (**CEU**). A score of 70 % is necessary to pass this course. We prefer if this exam is proctored. No intentional trick questions. If you should need any assistance, please email all concerns and the completed manual to info@tlch2o.com.

We would prefer that you utilize the enclosed answer sheet in the front, but if you are unable to do so, type out your own answer key. Please include your name and address on your answer key and make copy for yourself. You can e-mail or fax your Answer Key along with the Registration Form to TLC. **(S) Means answer may be plural or singular. Multiple Choice Section, One answer per question and please use the answer key.**

Disinfection Rules Chapter 1

What is in Water?

1. Water is primarily a liquid under standard conditions, which is not predicted from its relationship to other analogous hydrides of the oxygen family in the periodic table, which are liquids such as sulfuric acid.
A. True B. False
2. Water is the chemical substance with chemical formula H_2O : one molecule of water has two hydrogen atoms covalently bonded to a single oxygen atom.
A. True B. False
3. Water is a tasteless, odorless liquid at ambient temperature and pressure, and appears colorless in small quantities, although it has its own intrinsic very light blue hue. Ice also appears colorless, and water vapor is essentially invisible as a gas.
A. True B. False
4. Oxygen attracts protons much more strongly than hydrogen, resulting in a net negative charge on the hydrogen atoms, and a net negative charge on the oxygen atom. The presence of a charge on each of these atoms gives each water molecule a net dipole moment.
A. True B. False
5. Electrical attraction between water molecules due to this dipole pulls individual molecules closer together, making it more difficult to separate the molecules and therefore raising the boiling point.
A. True B. False

Disinfection Rule Review

6. Chlorine is the most widely used water disinfectant due to its effectiveness and cost. Using chlorine as a drinking water disinfectant has prevented millions of water borne diseases, such as typhoid, cholera, dysentery, and diarrhea. Most states require community water systems to use chlorination.

A. True B. False

7. Chlorine and chlorine-based compounds (halogens) react with organics in water causing the chlorine atom to substitute other atoms resulting in?

A. Chlorine D. Halogenated by-products
B. Organic sulfide(s) E. HOCl
C. Calcium carbonate F. None of the Above

8. Oxidation reactions, where chlorine oxidizes _____ present in water.

A. Carbon D. Chlorine and chlorine-based compounds (halogens)
B. Surface water E. Secondary by-products
C. Compounds F. None of the Above

9. Which of the following rules requires systems using public water supplies from either surface water or groundwater under the direct influence of surface water to disinfect?

A. TTHM and HAA5 Rule D. Disinfection byproducts (DBPs) Rule
B. DBP MCLs Rule E. Surface Water Treatment Rule (SWTR)
C. A community water system (CWS) F. None of the Above

10. At this time, an MCL is set for only _____, and proposed for additional disinfection byproducts.

A. TTHM and HAA5 Rule D. Disinfection byproducts (DBPs) Rule
B. DBP MCLs Rule E. Total Trihalomethanes
C. A community water system (CWS) F. None of the Above

11. Which of the following rules apply to all community and non-community water systems using a disinfectant such as chlorine, chloramines, ozone and chlorine dioxide?

A. TTHM and HAA5 Rule D. Disinfection byproducts (DBPs) Rule
B. DBP MCLs Rule E. Disinfectants and Disinfection Byproducts (DBP)
C. A community water system F. None of the Above

12. The Long Term 2 Enhanced Surface Water Treatment Rule (LT2) rule applies to all water systems using _____ under the influence of a surface water, as well as groundwater/surface water blends.

A. Surface water, groundwater D. Disinfection byproducts (DBPs) Rule
B. DBP MCLs Rule E. Total Trihalomethanes
C. A community water system (CWS) F. None of the Above

13. Which of the following rules began in 2006 with the characterization of raw water Cryptosporidium and E. coli levels. Systems serving <10,000 monitor for E. coli only every two weeks for one year. Compliance with the LT2 requirements began in April 2013?

- A. DBPs requirements
- B. Disinfectants requirements
- C. SDWA in 1996
- D. Stage 1 Disinfectant and Disinfection Byproduct Rule
- E. The LT2 requirements
- F. None of the Above

Public Health Concerns

14. According to the text, because of the large population exposed, health risks associated with _____, even if small, need to be taken seriously.

- A. DBPs
- B. Chlorine and chloramine
- C. Stage 2 DBPR
- D. Classes of DBPs
- E. Ultraviolet light
- F. None of the Above

15. Which of the following rules has been highly effective in protecting public health and has evolved to respond to new and emerging threats to safe drinking water?

- A. Stage 2 DBPR
- B. DBP exposure
- C. The Stage 2 DBP rule
- D. Long Term 2 Enhanced Surface Water Treatment Rule
- E. Safe Drinking Water Act (SDWA)
- F. None of the Above

16. _____ is one of the major public health advances in the 20th century..

- A. Major public health advances
- B. The Stage 2 DBPR
- C. Disinfection of drinking water
- D. Amendments to the SDWA in 1996
- E. Primary or residual disinfectant
- F. None of the Above

Microbial Regulations

17. One of the key regulations developed and implemented by the United States Environmental Protection Agency (USEPA) to counter pathogens in drinking water is the Surface Water Treatment Rule requires that a public water system, using surface water (or ground water under the direct influence of surface water) as its source, have sufficient treatment to reduce the source water concentration of Giardia and viruses by at least 99.9% and 99.99%, respectively.

- A. True
- B. False

18. Which rule specifies treatment criteria to assure that these performance requirements are met; they include turbidity limits, disinfectant residual, and disinfectant contact time conditions?

- A. Long Term 1 Rule
- B. Maximum Contaminant Level Goal (MCLG)
- C. Stage 1 Byproducts Rule
- D. Surface Water Treatment Rule
- E. Interim Enhanced Surface Water
- F. None of the Above

19. _____ was established to maintain control of pathogens while systems lower disinfection byproduct levels to comply with the Stage 1 Disinfectants/Disinfection Byproducts Rule and to control Cryptosporidium?

- A. Long Term 1 Enhanced Surface Water Treatment Rule
- B. Maximum Contaminant Level Goal (MCLG)
- C. Stage 1 Disinfectants/Disinfection Byproducts Rule
- D. Surface Water Treatment Rule
- E. Interim Enhanced Surface Water Treatment Rule
- F. None of the Above

20. The EPA established a MCL of 0.0010 for all public water systems and a 99% removal requirement for Cryptosporidium in filtered public water systems that serve at least 100,000 people. The new rule will tighten turbidity standards by December 2001.

- A. True B. False

21. Color is an indicator of the physical removal of particulates, including pathogens.

- A. True B. False

22. Which rule improves physical removal of Cryptosporidium, and to maintain control of pathogens?

- A. Long Term 1 Enhanced Surface Water Treatment Rule
B. Maximum Contaminant Level Goal (MCLG)
C. Stage 1 Disinfectants/Disinfection Byproducts Rule
D. Surface Water Treatment Rule
E. Interim Enhanced Surface Water Treatment Rule
F. None of the Above

Bromate

23. Fill in the missing information in order. _____ is a chemical that is formed when _____ used to disinfect drinking water reacts with naturally occurring _____ found in source water.

- A. Bromate, Ozone, Chlorite D. Hydrogen sulfide, Water, Nitrogen
B. Bromide, Bromate, Ozone E. Bromate, Ozone, Bromide
C. Bromate, Bromate, Bromate F. None of the Above

24. What is the annual average for Bromate that was established during the Stage 1 Disinfectants/Disinfection Byproducts Rule?

- A. 1 part per billion D. 10 parts per million
B. 10 parts per billion E. 500 parts per million
C. 100 parts per billion F. None of the Above

Chlorite

25. According to the Stage 1 Disinfectants/Disinfection Byproducts Rule, what is the monthly average level of chlorite in drinking water?

- A. 1 part per million D. 10 parts per million
B. 10 parts per billion E. 500 parts per million
C. 100 parts per billion F. None of the Above

Introduction to Chlorine (DDBP)

26. These term means that chlorine is present as Cl, HOCl, and OCl⁻ is called _____, and that which is bound but still effective is _____.

- A. Free available chlorine and Total D. Free available chlorine and Combined Chlorine
B. Free and Residual E. Combined chlorine and Readily available
C. Break point and Free F. None of the Above

27. Chloramines are formed by reactions with?

- A. Acid and Cl₂ D. Folic Acid and Cl₂
B. Ammonia and Cl₂ E. THMs and Haploidic acid
C. THMS and Cl₂ F. None of the Above

28. While testing chlorine disinfection process, you will need to understand one especially important feature is the ease of overdosing to create a "missing term" concentration.

- A. Free available chlorine and Total
- B. Residual
- C. Break point and Free
- D. Free available chlorine and Combined Chlorine
- E. Combined chlorine and readily available
- F. None of the Above

29. According to the text, this type of chlorine residual concentration residual is from 0.1 to 0.5 ppm.

- A. Free available chlorine and Total
- B. Residual
- C. Break point and Free
- D. Free available
- E. Combined chlorine and readily available
- F. None of the Above

30. Typically this chlorine residual is 2 ppm for _____.

- A. Free available chlorine and Total
- B. Residual
- C. Break point and Free
- D. Combined Chlorine
- E. Combined chlorine and readily available
- F. None of the Above

Chlorine By-Products

31. The most common chlorination by-products found in U.S. drinking water supplies are?

- A. Chlorate and Chlorite
- B. CO₂ and H₂SO₄
- C. Trihalomethanes (THMs)
- D. Ammonia and THMS
- E. Chloramines
- F. None of the Above

The Principal Trihalomethanes are:

32. Chloroform, bromodichloromethane, chlorodibromomethane, and bromoform. Other less common chlorination by-products include the haloacetic acids and haloacetonitriles. The amount of THMs formed in drinking water can be influenced by a number of factors, including the season and the source of the water.

- A. True
- B. False

33. THM concentrations are generally higher in winter than in summer, because concentrations of natural organic matter are greater and more chlorine is required to disinfect at colder temperatures.

- A. True
- B. False

34. THM levels are also low when wells or large lakes are used as the drinking water source, because organic matter concentrations are generally low in these sources. The opposite — high organic matter concentrations and high THM levels — is true when rivers or other surface waters are used as the source of the drinking water.

- A. True
- B. False

Health Effects

35. The available studies on health effects do not provide conclusive proof of a relationship between exposure to THMs and cancer or reproductive effects, but indicate the need for further research to confirm their results and to assess the potential health effects of chlorination by-products other than THMs.

- A. True
- B. False

Risks and Benefits of Chlorine

36. Many cities utilize the use of ozone to disinfect their source water and to reduce formation of this parameter?

- A. Chlorate and Chlorite
- B. CO₂ and H₂SO₄
- C. Trihalomethanes (THMs)
- D. Ammonia and THMS
- E. Chloramines
- F. None of the Above

37. _____ is a highly effective disinfectant; it breaks down quickly, so that small amounts of _____ or other disinfectants must be added to the water to ensure continued disinfection as the water is piped to the consumer's tap.

- A. Ozone, Chlorine
- B. UV, Chlorine
- C. Chlorite, Chlorine
- D. Chlorine Dioxide, Chlorine
- E. Chloramines, Chlorine
- F. None of the Above

38. Modifying water treatment facilities to use _____ can be expensive, and _____ treatment can create other undesirable by-products that may be harmful to health if they are not controlled (e.g., bromate).

- A. Ozone
- B. UV
- C. Chlorite
- D. Chlorine Dioxide
- E. Chloramines
- F. None of the Above

39. _____ is a weaker disinfectant than chlorine, especially against viruses and protozoa; they are very persistent and, as such, can be useful for preventing re-growth of microbial pathogens in drinking water distribution systems.

- A. Ozone
- B. UV
- C. Chlorite
- D. Chlorine Dioxide
- E. Chloramines
- F. None of the Above

40. Chlorine dioxide can be an effective disinfectant, but it forms?

- A. Chlorate and Chlorite
- B. CO₂ and H₂SO₄
- C. THMS
- D. Ammonia and THMS
- E. Chloramines
- F. None of the Above

41. It is extremely important that water treatment plants ensure that methods used to control chlorination by-products do not compromise the effectiveness of water disinfection.

- A. True
- B. False

Chloramines

42. What are chemical compounds formed by combining a specific ratio of chlorine and ammonia in water?

- A. Chlorine dioxide
- B. Bromate
- C. Chloramines
- D. Trihalomethanes, haloacetic acids, bromate, and chlorite
- E. Disinfection byproducts
- F. None of the Above

43. _____ provides a durable residual, and are often used as a secondary disinfectant for long distribution lines and where free chlorine demand is high.

- A. Chlorine dioxide
- B. Bromate
- C. Chloramines
- D. Trihalomethanes, haloacetic acids, bromate, and chlorite
- E. Disinfection byproducts
- F. None of the Above

44. This term represents a compound that may also be used instead of chlorine in order to reduce chlorinated byproduct formation and to remove some taste and odor problems.
- A. Chlorine dioxide
 - B. Bromate
 - C. Chloramines
 - D. Trihalomethanes, haloacetic acids, bromate, and chlorite
 - E. Disinfection byproducts
 - F. None of the Above

Chlorine Dioxide

45. Which term represents a compound that may be generated on-site at water treatment facilities? In most generators, sodium chlorite and elemental chlorine are mixed in solution, which almost instantaneously forms?

- A. Chlorine dioxide (ClO₂)
- B. Bromate
- C. Chloramine
- D. Ozone
- E. Disinfection compounds
- F. None of the Above

46. In solution, Chlorine dioxide it is a dissolved gas, which makes it largely unaffected by pH but volatile and relatively easily stripped from?

- A. Chlorine
- B. Sodium hypochlorite
- C. Chlorine dioxide
- D. Chlorine gas
- E. Solution
- F. None of the Above

47. Which of the following terms does chlorine dioxide produce?

- A. CO₂ and H₂SO₄
- B. Sodium hypochlorite
- C. Bromate
- D. Trihalomethanes, haloacetic acids, bromate, and chlorite
- E. A residual it is only rarely used for this purpose
- F. None of the Above

Factors in Chlorine Disinfection: Concentration and Contact Time

48. Which of the following terms is multiplied by minimum contact time (minutes), offer water operator's guidance in computing an effective combination of chlorine concentration and chlorine contact time required to achieve disinfection of water at a given temperature?

- A. CXT concept
- B. CXT values
- C. CXT formula
- D. Pound per day
- E. Contact concept
- F. None of the Above

49. _____ demonstrates that if an operator chooses to decrease the chlorine concentration, the required contact time must be lengthened.

- A. CXT concept
- B. CXT values
- C. CXT formula
- D. Contact concept
- E. Pound per day
- F. None of the Above

50. When free available chlorine residuals are desired, the characteristics of the water will determine how this will be accomplished. This may have to be considered:

If the water contains no ammonia or this missing term, any application of chlorine will yield a free residual once it has reacted with any bacteria, virus and other microorganisms present in the water.

- A. Other nitrogen compounds
- B. More chlorine
- C. Taste and odor
- D. Amount of chlorine
- E. Iron, manganese, organic matter
- F. None of the Above

51. If the water contains _____, it results in the formation of a combined residual, which must be destroyed by applying an excess of chlorine.

- A. Free available residual
- B. Free available chlorine
- C. Chloramines
- D. Additional chlorine
- E. Ammonia
- F. None of the Above

52. Breakpoint chlorination is the name of the process of adding chlorine to water until the chlorine demand has been satisfied.

- A. True
- B. False

53. This term means the amount of chlorine used up before a free available chlorine residual is produced.

- A. Free available residual
- B. Free available chlorine
- C. Chlorine demand
- D. No disinfection
- E. No residual
- F. None of the Above

54. What is the addition of chlorine that results in a chlorine residual that is directly proportional to the amount of chlorine added beyond the?

- A. Free available residual
- B. Free available chlorine
- C. Breakpoint
- D. No disinfection
- E. No residual
- F. None of the Above

Understanding Waterborne Viruses

55. Which of the following terms does treatment processes and watershed management strategies do not necessarily protect against?

- A. F-specific coliphages
- B. Indicators of fecal contamination
- C. Enteric viruses
- D. Viruses
- E. Viral infection
- F. None of the Above

56. Which is the current method for culturing this missing term under the ICR (U.S. Environmental Protection Agency, 1996c) is recognized as being difficult to implement.

- A. Protozoan pathogens
- B. Indicators of fecal contamination
- C. Enteric viruses
- D. Viruses
- E. Microorganisms
- F. None of the Above

57. _____ is found in environmental samples presumably come from warm-blooded animals or sewage?

- A. F-specific coliphages
- B. Indicators of fecal contamination
- C. Enteric viruses
- D. Viruses
- E. Microorganisms
- F. None of the Above

58. This term represents coliphage is representative of the survival and transport of?

- A. Enteric viruses
- B. F-specific coliphages
- C. Viruses
- D. Indicators of viral contamination
- E. Coliphages are bacteriophages
- F. None of the Above

Sampling Procedures

Streamwater Sample Collection

59. Consider that the spatial and temporal distribution of microorganisms in surface water can be as variable as the distribution of this missing term because microorganisms are commonly associated with solid particles.

- A. Suspended sediment
- B. Indicators of fecal contamination
- C. Enteric viruses
- D. Viruses
- E. Microorganisms
- F. None of the Above

60. The standard samplers used in by the majority of samplers can be used to collect streamwater samples for bacterial and viral indicators, this missing term providing that the equipment coming in contact with the water is properly cleaned and sterilized.

- A. Cryptosporidium, and Giardia
- B. Indicator organisms
- C. Cholera, polio, typhoid, hepatitis
- D. Cryptosporidium
- E. Giardia
- F. None of the Above

Cryptosporidium and Giardia Analysis

61. Special sterilization procedures are needed for equipment used in the collection of samples for _____.

- A. Total Organisms
- B. Indicator bugs
- C. Cholera, polio, typhoid, hepatitis
- D. Oocysts
- E. Cryptosporidium and Giardia
- F. None of the Above

Understanding Bacteriophage

62. Bacteriophages may have a lytic cycle or a lysogenic cycle, such as the T4 phage, this missing term are broken open (lysed) and destroyed after immediate replication of the virion.

- A. Lysogenic cycle
- B. Bacterial cells
- C. Vibrio cholerae
- D. Phage virions
- E. Myovirus bacteriophages
- F. None of the Above

63. _____ is an example of a virus that remains dormant until host conditions deteriorate, perhaps due to depletion of nutrients; then it becomes active?

- A. Lysogenic cycle
- B. Endogenous phages
- C. Vibrio cholerae
- D. Phage virions
- E. Myovirus bacteriophages
- F. None of the Above

64. Which of the following is an example of a bacteriophage known to follow the lysogenic cycle and the lytic cycle is the?

- A. Podoviruses
- B. Phage's host range
- C. Myovirus bacteriophages
- D. Phage lambda of E. coli
- E. Viral genome
- F. None of the Above

65. Which of the following terms is an example is the conversion of a harmless strain of a phage that can cause cholera?

- A. Lysogenic cycle
- B. The virus
- C. Vibrio cholerae
- D. Phage virions
- E. Myovirus bacteriophages
- F. None of the Above

Understanding Disinfection

Wastewater Disinfection

66. According to the text, there are a number of chemicals and processes that will this missing term, but none are universally applicable.

- A. Limit the effects of organic material
- B. Numerous alternative disinfection processes
- C. Residual level of disinfection
- D. Disinfect wastewater
- E. Limit the travel of pathogens
- F. None of the Above

Water Disinfection

67. Disinfection is usually the final stage in the _____ in order to limit the effects of organic material, suspended solids and other contaminants.

- A. Limit the effects of organic material
- B. Numerous alternative disinfection processes
- C. Residual level of disinfection
- D. Water treatment process
- E. Limit the travel of pathogens
- F. None of the Above

68. _____ expresses that this is less widely used in small and very small water treatment systems, including chlorine dioxide, potassium permanganate, chloramines and peroxone (ozone/hydrogen peroxide)?

- A. Limit the effects of organic material
- B. Numerous alternative disinfection processes
- C. Residual level of disinfection
- D. Additional killing mechanism
- E. Pathogens
- F. None of the Above

69. According to the text, surface waters have been the focal point of _____ since their inception, as groundwaters (like wells) have been historically considered to be free of microbiological contamination.

- A. Chlorates are powerful oxidizers
- B. Adverse health effects
- C. Water disinfection regulations
- D. Microbiological contamination
- E. Sodium chloride
- F. None of the Above

Chlorate Ion

70. Which of the following terms is predicted by VSEPR, about chlorate anions?

- A. Acid/base balance
- B. Stable perchlorates
- C. Formula ClO₃⁻
- D. Trigonal pyramidal structures
- E. Chemical formula CaCl₂
- F. None of the Above

71. According to the text, _____ and should be kept away from organics or easily oxidized materials?

- A. Chlorates are powerful oxidizers
- B. Adverse health effects
- C. Formula ClO₃⁻
- D. Microbiological contamination
- E. Sodium chloride
- F. None of the Above

72. Chlorates were once widely used in _____ for this reason, though their use has fallen due to their instability. Most pyrotechnic applications which formerly used chlorates in the past now use the more stable perchlorates instead.

- A. Acid/base balance
- B. Stable perchlorates
- C. Formula ClO₃⁻
- D. Pyrotechnics
- E. Chemical formula CaCl₂
- F. None of the Above

Chloride Ion

73. The chloride ion is formed when the _____, a halogen, gains an electron to form an anion (negatively-charged ion) Cl⁻.

- A. Chlorates are powerful oxidizers
- B. Adverse health effects
- C. Element chlorine
- D. Microbiological contamination
- E. Sodium chloride
- F. None of the Above

74. The salts of hydrochloric acid contain chloride ions and can also be called chlorides. The chloride ion, and its salts such as sodium chloride, _____.

- A. Acid/base balance
- B. The stable perchlorates
- C. The formula ClO⁻³
- D. Are very soluble in water
- E. The chemical formula CaCl₂
- F. None of the Above

75. Methyl chloride, more commonly called chloromethane, (CH₃Cl) is _____, which does not contain a chloride ion.

- A. Chlorates are powerful oxidizers
- B. Adverse health effects
- C. The chloride ion
- D. An organic covalently bonded compound
- E. Sodium chloride
- F. None of the Above

76. Other salts such as calcium chloride, magnesium chloride, potassium chloride have varied uses ranging from medical treatments to?

- A. Chemical formula CaCl₂
- B. Cement formation
- C. Chlorite ion is ClO₂⁻.
- D. Corresponding anions Cl⁻, ClO⁻, ClO₂⁻, ClO₃⁻, or ClO₄⁻
- E. Chlorine dioxide
- F. None of the Above

77. Which of the following compounds is an example is table salt?

- A. CaCl₂
- B. NaCl
- C. ClO₂⁻.
- D. Cl⁻, ClO⁻, ClO₂⁻, ClO₃⁻, or ClO₄⁻
- E. NaCl₂
- F. None of the Above

78. _____ is also the prosthetic group present in the amylase enzyme. Another example is calcium chloride with the chemical formula CaCl₂.

- A. Chemical formula CaCl₂
- B. A chloride ion
- C. Chlorite ion is ClO₂⁻.
- D. Corresponding anions Cl⁻, ClO⁻, ClO₂⁻, ClO₃⁻, or ClO₄⁻
- E. Chlorine dioxide
- F. None of the Above

79. Which of the following compounds is used for maintaining unpaved roads and for sanite fortifying roadbases for new construction?

- A. Chemical formula CaCl₂
- B. Calcium chloride
- C. Chlorite ion is ClO₂⁻.
- D. Corresponding anions Cl⁻, ClO⁻, ClO₂⁻, ClO₃⁻, or ClO₄⁻
- E. Chlorine dioxide
- F. None of the Above

80. Which of the following compounds are a closely monitored constituent of the mud system?

- A. Chemical formula CaCl₂
- B. Chloride
- C. Chlorite ion is ClO₂⁻.
- D. Chlorides
- E. Chlorine dioxide
- F. None of the Above

81. _____ is also a useful and reliable chemical indicator of river / groundwater fecal contamination, as chloride is a non-reactive solute and ubiquitous to sewage & potable water.

- A. Chemical formula CaCl_2
- B. Chloride
- C. Chlorite ion is ClO_2^- .
- D. Corresponding anions Cl^- , ClO^- , ClO_2^- , ClO_3^- , or ClO_4^-
- E. Chlorine dioxide
- F. None of the Above

Chlorite Ion

82. The chlorite ion is _____.

- A. Chemical formula CaCl_2
- B. Chloride
- C. ClO_2^-
- D. Corresponding anions Cl^- , ClO^- , ClO_2^- , ClO_3^- , or ClO_4^-
- E. Chlorine dioxide
- F. None of the Above

83. Chlorine can assume oxidation states of -1, +1, +3, +5, or +7 within the corresponding anions Cl^- , ClO^- , ClO_2^- , ClO_3^- , or ClO_4^- , known commonly and respectively as _____.

- A. Chemical formula CaCl_2
- B. Chloride
- C. Chlorite ion is ClO_2^- .
- D. Chloride, hypochlorite, chlorite, chlorate, and perchlorate
- E. Chlorine dioxide
- F. None of the Above

84. An additional oxidation state of +4 is seen in the neutral compound _____, which has a similar structure to chlorite ClO_2^- (oxidation state +3) and the cation chloryl (ClO_2^+) (oxidation state +5).

- A. Chemical formula CaCl_2
- B. Chloride
- C. Chlorite ion is ClO_2^- .
- D. Corresponding anions Cl^- , ClO^- , ClO_2^- , ClO_3^- , or ClO_4^-
- E. Chlorine dioxide ClO_2
- F. None of the Above

Chlorine Dioxide

85. Chlorine dioxide is a chemical compound with the formula?

- A. Chemical formula CaCl_2
- B. Chloride
- C. Chlorite ion is ClO_2^- .
- D. Corresponding anions Cl^- , ClO^- , ClO_2^- , ClO_3^- , or ClO_4^-
- E. ClO_2
- F. None of the Above

Haloacetic Acids

86. What type of substances are haloacetic acids in which a halogen atom takes the place of a hydrogen atom in acetic acid?

- A. An anti-bonding orbital
- B. A single halogen
- C. Hypochlorite compounds
- D. Carboxylic acids
- E. Calcium hypochlorite
- F. None of the Above

Contaminants in Drinking Water

87. _____ expresses an exposure to such substances in drinking water has been associated with a number of health outcomes by epidemiological studies, although the putative agent in such studies has not been identified.

- A. An anti-bonding orbital
- B. A single halogen
- C. Hypochlorite compounds
- D. Disinfection by-products
- E. Calcium hypochlorite
- F. None of the Above

Hypochlorites

88. The same residuals are obtained as with gas chlorine, but the effect on the _____ of the treated water is different.
- A. High-test calcium hypochlorite(s)
 - B. Calcium hypochlorite tablets
 - C. Hypochlorous acid
 - D. Negative charge
 - E. pH
 - F. None of the Above

Disinfection Byproducts

89. Which term represents when disinfectants used in water treatment plants react with bromide and/or natural organic matter (i.e., decaying vegetation) present in the source water?
- A. Disinfection byproducts
 - B. Other disinfectants
 - C. Naturally occurring bromide
 - D. Occurring organic and inorganic matter in water
 - E. Most prevalent THM
 - F. None of the Above
90. _____ which regulations have been established have been identified in drinking water, including trihalomethanes, haloacetic acids, bromate, and chlorite?
- A. Chlorine dioxide
 - B. HAA5
 - C. Trihalomethanes
 - D. Trihalomethanes, haloacetic acids, bromate, and chlorite
 - E. Disinfection byproducts
 - F. None of the Above

Trihalomethanes (THM)

91. This term represents a group of four chemicals that are formed along with other disinfection byproducts when chlorine or other disinfectants used to control microbial contaminants in drinking water react with naturally occurring organic and inorganic matter in water.
- A. Disinfection byproducts
 - B. Other disinfectants
 - C. Naturally occurring bromide
 - D. Occurring organic and inorganic matter in water
 - E. Trihalomethanes (THM)
 - F. None of the Above
92. _____ represents chloroform, bromodichloromethane, dibromochloromethane, and bromoform?
- A. Chlorine dioxide
 - B. HAA5
 - C. Trihalomethanes
 - D. Trihalomethanes, haloacetic acids, bromate, and chlorite
 - E. Chloroform
 - F. None of the Above

Haloacetic Acids (HAA5)

93. This term represents substances in drinking water that reacts with naturally occurring organic and inorganic matter in water.
- A. Disinfection byproducts
 - B. Other disinfectants
 - C. Naturally occurring bromide
 - D. Occurring organic and inorganic matter in water
 - E. Microbial contaminants
 - F. None of the Above
94. _____ represents monochloroacetic acid, dichloroacetic acid, trichloroacetic acid, monobromoacetic acid, and dibromoacetic acid.
- A. Chlorine dioxide
 - B. HAA5
 - C. Trihalomethanes
 - D. Trihalomethanes, haloacetic acids, bromate, and chlorite
 - E. Chloroform
 - F. None of the Above

95. Bromate is a chemical that is formed when _____ is used to disinfect drinking water?

- A. Disinfection byproducts
- B. Other disinfectants
- C. Naturally occurring bromide
- D. Occurring organic and inorganic matter in water
- E. Ozone
- F. None of the Above

96. This term represents a byproduct formed when chlorine dioxide is used to disinfect water.

- A. Chlorite
- B. HAA5
- C. Trihalomethanes
- D. Trihalomethanes, haloacetic acids, bromate, and chlorite
- E. Chloroform
- F. None of the Above

Chloroform

97. Which is typically the most prevalent THM measured in chlorinated water, is probably the most thoroughly studied disinfection byproduct?

- A. Disinfection byproducts
- B. Other disinfectants
- C. Naturally occurring bromide
- D. Occurring organic and inorganic matter in water
- E. Chloroform
- F. None of the Above

Sodium Chlorate

98. Sodium chlorate can also be synthesized by passing _____ into a hot sodium hydroxide solution. It is then purified by crystallization.

- A. Chlorate
- B. Oxygen
- C. Chlorine gas
- D. Sodium metaborate or ammonium phosphates
- E. The free acid, chlorous acid, HClO_2
- F. None of the Above

Sample Procedures

99. Samples need to be kept on ice and shipped to a central laboratory for analysis of coliphage, *C. perfringens*, *Cryptosporidium*, *Giardia*, and enteric viruses by the current analytical methods. The single-agar layer (SAL), direct plating method with induction of streptomycin and ampicillin is recommended for detection of somatic and F-specific coliphage in streamwater samples.

- A. True
- B. False

Disinfection Byproduct Regulations Summary

100. Water system managers may also consider switching from chlorine to alternative disinfectants to reduce formation of THMs and HAAs.

- A. True
- B. False

Halogen Section Chapter 2

Halides

101. Halide ions combined with single hydrogen atoms form the hydrohalic acids (i.e., HF, HCl, HBr, HI), a series of particularly strong acids, one being?

- A. HCl
- B. HOCl
- C. Hydrastatic acid
- D. Chlorine gas
- E. The hypochlorite ion (OCI-)
- F. None of the Above

102. Many synthetic organic compounds such as plastic polymers, and a few natural ones, contain halogen atoms; these are known as halogenated compounds or?

- A. Organic halides
- B. Free radicals
- C. Diatomic Compound
- D. Many synthetic organic compounds
- E. Neither fluorine nor bromine
- F. None of the Above

103. What is the negative ion referred to as?

- A. Salts
- B. A halide proton
- C. A halide ion
- D. Free radical
- E. Diatomic Compound
- F. None of the Above

104. Which of the following terms containing these ions are known as halides?

- A. Salts
- B. CXT values
- C. Primary disinfectant
- D. Many synthetic organic compounds
- E. Neither fluorine nor bromine
- F. None of the Above

Chlorine

105. This halogen is needed only in very small amounts for the production of thyroid hormones such as thyroxine?

- A. Chlorine
- B. Chlorine dioxide
- C. Iodine
- D. Halogen(s)
- E. Inhibitory transmitter GABA
- F. None of the Above

106. On the other hand, neither fluorine nor bromine are believed to be really essential for humans, although small amounts of _____ can make tooth enamel resistant to decay.

- A. Salts
- B. Iodine
- C. Chlorine
- D. Synthetic organic compounds
- E. Fluoride
- F. None of the Above

107. Which halogen is needed in relatively large amounts by humans?

- A. Chlorine
- B. Chlorine dioxide
- C. Iodine
- D. Halogen(s)
- E. Inhibitory transmitter GABA
- F. None of the Above

Halogens

108. All Halogens have 7 electrons in their outer shells, giving them an oxidation number of -1. The halogens exist, at room temperature, in all three states of matter:

- A. True
- B. False

pH Section

109. What is a substance that has the ability to reduce other substances and is said to be reductive in nature?

- A. Protons
- B. An electron donor
- C. Anti-matter
- D. Electrons
- E. Cations
- F. None of the Above

110. Pure water has a pH very close to?

- A. 5
- B. 6
- C. 7
- D. 7.7
- E. 7.5
- F. None of the Above

111. One definition of pH is that it is defined as the decimal logarithm of the reciprocal of the _____, a_{H^+} , in a solution.

- A. Hydrogen ion activity
- B. Ion-selective electrode(s)
- C. (Solvated) hydronium ion
- D. Brønsted–Lowry acid–base theory
- E. Acid-base behavior
- F. None of the Above

112. Because the pH scale is logarithmic, therefore pH is?

- A. Universal indicator
- B. A dimensionless quantity
- C. A Spectrophotometer
- D. Excess of Ion concentrations
- E. A set of non-linear equations
- F. None of the Above

113. Alkalinity is able to neutralize _____ and is measured in a quantitative capacity in an aqueous solution.

- A. Acid
- B. Base
- C. pH
- D. pH measurement(s)
- E. Bond formation
- F. None of the Above

114. For strong acids and bases no calculations are necessary except in extreme situations. The pH of a solution containing a weak acid requires the solution of a quadratic equation.

- A. True
- B. False

115. If the pH of a solution contains a weak base, this may require?

- A. The solution of a cubic equation
- B. The solution of a linear equation
- C. The solution of a squared equation
- D. A set of linear simultaneous equations
- E. A set of non-linear simultaneous equations
- F. None of the Above

116. While the general case requires the pH solution of?

- A. The solution of a cubic equation
- B. The solution of a linear equation
- C. The solution of a squared equation
- D. A set of linear simultaneous equations
- E. A set of non-linear simultaneous equations
- F. None of the Above

117. Because alkalinity is significant in many uses and treatments of natural waters and wastewaters, the measured values also may include contributions from _____ or other bases if these are present.

- A. Acids
- B. Light metals
- C. Rare earths
- D. Borates, phosphates, silicates
- E. Caustics
- F. None of the Above

118. Calculations are not necessary except in extreme situations for strong acids and bases. The pH of a solution containing a weak acid requires?

- A. The concentration value
- B. The solution of a quadratic equation
- C. The Spectrophotometer
- D. Visual comparison
- E. The solution of a cubic equation
- F. None of the Above

119. Since pH is a logarithmic scale, a difference of one pH unit is equivalent to a _____ difference in hydrogen ion concentration

- A. 1
- B. 2
- C. 5
- D. 10
- E. 100
- F. None of the Above

120. According to the manual, this key water measurement is used in the interpretation and control of water and wastewater treatment processes.

- A. Acid
- B. Alkalinity
- C. pH
- D. Chemical ion
- E. Hydrogen bond formation
- F. None of the Above

121. These compounds for all practical purposes are completely dissociated in water.

- A. Strong acids and bases
- B. Strong bases
- C. Chemical ions in chains
- D. Strong bases and weak acids
- E. Weak acids and weak bases
- F. None of the Above

122. Sodium hydroxide, NaOH, is an example of _____.

- A. Strong acid and base
- B. Strong base
- C. Weak base
- D. Strong base and weak acid
- E. Weak acids and weak bases
- F. None of the Above

123. According to the text, what is the pH of pure water at 50 °C?

- A. 7.7
- B. 8.0
- C. 9.0
- D. 6.55
- E. 7.00
- F. None of the Above

Hard Water Section

124. Water contains various amounts of _____, some of which impart a quality known as hardness. Consumers frequently complain about problems attributed to hard water, such as the formation of scale on cooking utensils and hot water heaters.

- A. Water hardness
- B. Carbonate hardness
- C. The calcium-magnesium distinction
- D. Calcium (Ca) and magnesium (Mg)
- E. Dissolved minerals
- F. None of the Above

Occurrence of Hard Water

125. Hard water is caused by soluble, divalent, _____, (positive ions having valence of 2). The principal chemicals that cause water hardness are calcium (Ca) and magnesium (Mg).

- A. Water hardness
- B. Metallic cations
- C. Carbon dioxide (CO₂)
- D. Calcium (Ca) and magnesium (Mg)
- E. Noncarbonate hardness
- F. None of the Above

Chlorine Section Chapter 3

Chlorine's Appearance and Odor (QA/QC)

126. Chlorine is a greenish-yellow gas it will condense to an amber liquid at approximately _____ F or at high pressures.

- A. 32 degrees
- B. - 100 degrees
- C. 129 degrees
- D. 29 degrees
- E. -29.2 degrees
- F. None of the Above

127. Prolonged exposures to chlorine gas may result in _____. Odor thresholds ranging from 0.08 to part per million (ppm) parts of air have been reported.

- A. Exposure to chlorine
- B. Odor thresholds
- C. A corrosive material
- D. Olfactory fatigue
- E. Moisture, steam, and water
- F. None of the Above

Reactivity

128. Contact between chlorine and arsenic, bismuth, boron, calcium, activated carbon, carbon disulfide, glycerol, hydrazine, iodine, methane, oxomonosilane, potassium, propylene, and silicon should be avoided.

- A. True
- B. False

129. Chlorine reacts with hydrogen sulfide and water to form this substance?

- A. Hydrogen sulfide
- B. Oxomonosilane
- C. Sodium Chloride
- D. Chlorinates
- E. Hydrochloric acid
- F. None of the Above

130. According to the text, chlorine is also incompatible with _____.

- A. Air
- B. Ammonia
- C. Sodium Chloride
- D. Hydrogen sulfide
- E. Moisture, steam, and water
- F. None of the Above

Flammability

131. When there is a fire that involves chlorine, the fire fight should be fought downwind from the minimum distance possible.

- A. True
- B. False

Chlorine Basics

132. By dropping a few drops of hydrochloric acid onto a piece of manganese dioxide, Steele had discovered?

- A. Halogens
- B. Ammonia
- C. Chlorine
- D. Manganese dioxide
- E. H₂SO₄
- F. None of the Above

133. English chemist Sir Humphrey Davy discovered that chlorine gas was actually an element. Until that time, people were convinced that the gas was a?

- A. Chlorine chemistry
- B. Compound of oxygen
- C. Compound of ammonia
- D. Economical germ-killers
- E. Theory
- F. None of the Above

What Happens to Chlorine When it Enters the Environment?

134. When chlorine is released to soil, chlorine will react with moisture forming?

- A. Free oxygen radicals
- B. Chlorine gas
- C. Hydrochloric acid
- D. A greenish-yellow, noncombustible gas
- E. Hypochlorous acid and hydrochloric acid
- F. None of the Above

135. When released to air, chlorine will react with water to form hypochlorous acid and hydrochloric acid, which are removed from the atmosphere by generation of free oxygen radicals.

- A. True
- B. False

136. Chlorine reacts with water to form hypochlorous acid and hydrochloric acid. The hypochlorous acid breaks down rapidly. The hydrochloric acid also breaks down; its breakdown products will raise the pH of the water (makes it more basic).

- A. True
- B. False

Disinfectant Qualities

137. Chlorine is so important in poultry processing that the US Department of Agriculture requires an almost constant chlorine rinse for much of the cutting equipment. In fact, no proven economical alternative to chlorine disinfection exists for use in Meat and poultry processing facilities.

- A. True
- B. False

Properties

138. Because it is highly reactive, chlorine is usually found in nature bound with other elements like sodium, potassium, and magnesium.

- A. True
- B. False

139. Chemical elements have their own set of unique properties and chlorine is known as _____--so reactive, in fact, that it is usually found combined with other elements in the form of compounds.

- A. Synthesizing organic compounds
- B. A very reactive element
- C. Chlorine compounds
- D. Organic compounds
- E. One of the most abundant chemical elements
- F. None of the Above

140. Which substance is capable of removing a wide variety of disease-causing germs from drinking water and wastewater as well as from hospital and food production surfaces?

- A. Inorganic disinfectant
- B. Chlorine-based disinfectants
- C. Ancient seawater
- D. Useful chemical elements
- E. Organic compounds
- F. None of the Above

141. Various states of chlorine includes when chlorine is isolated as a free element, chlorine is a greenish yellow gas, which is _____.

- A. 2.5 times heavier than water
- B. 2.5 times lighter than air
- C. 10 times heavier than air
- D. 2.5 times heavier than air
- E. 25 times heavier than air
- F. None of the Above

Released From the Salt of the Earth

142. According to the text, because of the slow evaporation of _____, Chlorine is produced industrially from the compound sodium chloride.

- A. Water
- B. Brine
- C. Ancient seawater
- D. Useful chemical elements
- E. Organic compounds
- F. None of the Above

Chlorine's Effectiveness

143. The amount of chlorine required to achieve disinfection and that reacts with the other chemicals is the?

- A. Chlorine residual
- B. Color change
- C. Chlorine demand
- D. Total
- E. Free chlorine residual
- F. None of the Above

144. Which of the following terms to disinfect decreases, as the concentration of the chlorine increases?

- A. pH increases
- B. Chlorine level and water quality
- C. Free chlorine residual
- D. Required contact time
- E. Not available for disinfection
- F. None of the Above

145. Chlorination is less effective in _____.

- A. Clear water
- B. Color change
- C. Warm temps
- D. Day time
- E. Cloudy (turbid) water
- F. None of the Above

146. By adding a little more chlorine to what is already sufficient, this action will generally result in _____ that can be measured easily.

- A. pH increases
- B. Chlorine level and water quality
- C. Chlorine demand
- D. Required contact time
- E. A free chlorine residual
- F. None of the Above

Chemistry of Chlorination

147. The hypochlorite ion is a much weaker disinfecting agent than Hypochlorous acid, about 100 times less effective.

- A. True
- B. False

148. According to the text, pH and temperature affect the ratio of hypochlorous acid to hypochlorite ions. As the temperature is decreased, the _____ increases.

- A. Reduction Ratio
- B. CT actual
- C. Free chlorine residual
- D. "CT" disinfection concept
- E. Ratio of hypochlorous acid
- F. None of the Above

149. Under normal water conditions, hypochlorous acid will also chemically react and break down into the hypochlorite ion.

- A. True
- B. False

150. Although the ratio of _____ is greater at lower temperatures, pathogenic organisms are actually harder to kill.

- A. Hypochlorous acid
- B. The amount of chlorine
- C. Chlorine Demand
- D. Total chlorine
- E. pH value and temperature
- F. None of the Above

151. If all other things were equal, _____ and a lower pH are more conducive to chlorine disinfection.

- A. Lower pH
- B. Hypochlorous acid
- C. Higher water temperatures
- D. Lower water temperature
- E. The hypochlorite ion
- F. None of the Above

152. The disassociation of chlorine gas

(OCI⁻): HOCl → H⁺ + OCl⁻ Also expressed HOCl → H⁺ + OCl⁻
(hypochlorous acid) (hydrogen) (hypochlorite ion)

- A. True
- B. False

153. All three forms of chlorine produce Sodium hypochlorite when added to water.

- A. True
- B. False

Types of Residual

154. Which of the following terms is all chlorine that is available for disinfection?

- A. Chlorine residual
- B. Chlorine demand
- C. Free chlorine
- D. Break-point chlorination
- E. Total chlorine
- F. None of the Above

155. Total chlorine residual = free + _____.

- A. Chlorine residual
- B. Chlorine demand
- C. Free chlorine
- D. Combined chlorine residual
- E. Total chlorine residual
- F. None of the Above

156. Which of the following terms is where the chlorine demand has been satisfied, and any additional chlorine will be considered free chlorine?

- A. Chlorine residual
- B. Chlorine demand
- C. Free chlorine
- D. Break-point chlorination
- E. Total chlorine residual
- F. None of the Above

Residual Concentration/Contact Time (CT) Requirements

157. Since monitoring for very low levels of pathogens in treated water is analytically very difficult, utilizing the _____ is recommended to demonstrate satisfactory treatment.

- A. Free chlorine
- B. Total residual
- C. Free chlorine residual
- D. "CT" disinfection concept
- E. T10 of the process unit
- F. None of the Above

158. Fill in the blank _____ = Concentration (mg/L) x Time (minutes)

- A. CT
- B. The amount of chlorine
- C. Chlorine Demand
- D. Total chlorine
- E. pH value and temperature
- F. None of the Above

159. The effective reduction in pathogens can be calculated by reference to standard tables of required?

- A. Free chlorine
- B. Total residual
- C. Free chlorine residual
- D. "CT" s
- E. T10 of the process unit
- F. None of the Above

160. The CT concept as developed by the United States Environmental Protection Agency (uses the combination of disinfectant residual concentration (mg/L) and the effective disinfection contact time (in minutes) to measure effective pathogen reduction.

- A. True
- B. False

Calculation and Reporting of CT Data

161. The _____ must be greater than 1.0 to be acceptable.

- A. Reduction Ratio
- B. CT actual
- C. Free chlorine residual
- D. "CT" disinfection concept
- E. T10 of the process unit
- F. None of the Above

162. You can also calculate and record actual log reductions. Reduction Ratio = CT actual divide by?

- A. Reduction Ratio
- B. CT
- C. Free chlorine residual
- D. "CT" disinfection concept
- E. CT required
- F. None of the Above

Chlorine Review

163. _____ describes the minimum amount of Chlorine needed to react in a water purification system; used as a monitoring measurement by system operators?

- A. Chlorine Demand
- B. Liquid
- C. Total chlorine
- D. Monitoring measurement
- E. Ammonia or organic amines
- F. None of the Above

164. Operator may add _____ to chlorinated public water supplies to provide inorganic chloramines.

- A. Combined chlorine
- B. Liquid
- C. Total chlorine
- D. Ammonia
- E. Organic amines
- F. None of the Above

165. Which term describes the concentration of residual chlorine in water present as dissolved gas (Cl_2), hypochlorous acid (HOCl), and/or hypochlorite ion (OCl^-)?

- A. Chlorine Demand
- B. Chlorine total
- C. Free Chlorine
- D. Total chlorine
- E. Residual chlorine
- F. None of the Above

166. Which term describes the concentration of chlorine in the water after the chlorine demand has been satisfied?

- A. Chlorine Residual
- B. Chlorine Demand
- C. Combined
- D. Total chlorine
- E. Residual chlorine
- F. None of the Above

167. Which term describes the amount of chlorine used up in a water purification system; used as a monitoring measurement by system operators?

- A. Chlorine Residual
- B. Chlorine Demand
- C. Combined Chlorine Residual
- D. Total chlorine
- E. Residual chlorine
- F. None of the Above

168. _____ the residual chlorine existing in water in chemical combination with ammonia or organic amines which can be found in natural or polluted waters.

- A. Chlorine Residual
- B. Chlorine Demand
- C. Combined Chlorine Residual
- D. Total chlorine
- E. Residual chlorine
- F. None of the Above

169. Which of the following terms of at least 1.0 mg/L should be maintained in the clear well or distribution reservoir immediately downstream from the point of post-chlorination and .2 mg/L in the distribution system to guard against backflow.

- A. Chlorine Demand
- B. Chlorine total
- C. Free chlorine residual
- D. Total chlorine
- E. Residual chlorine
- F. None of the Above

170. _____ describes the total of free residual and combined residual chlorine in a water purification system; and used as a monitoring measurement by system operators.

- A. Chlorine Demand
- B. Chlorine total
- C. Total Chlorine Residual
- D. Total combined chlorine
- E. Residual chlorine
- F. None of the Above

Common Terms

171. What is the best term that describes chlorine addition of chlorine at the plant headworks or prior to other water treatment or groundwater production processes and mainly used for disinfection and control of tastes, odors, and aquatic growth?

- A. Chlorination
- B. Post-chlorination
- C. Chlorine Demand
- D. Demand
- E. Pre-chlorination
- F. None of the Above

172. _____ describes the sum of free and combined chlorine.

- A. Organic amine(s)
- B. Disinfection
- C. Free chlorine
- D. Breakpoint chlorination
- E. Total Chlorine
- F. None of the Above

173. When chlorinating most potable water supplies, total chlorine is essentially equal to _____ since the concentration of ammonia or organic nitrogen compounds (needed to form combined chlorine) will be very low.

- A. Chlorination
- B. The amount of chlorine
- C. Chlorine Demand
- D. Total chlorine
- E. Free chlorine
- F. None of the Above

174. What term best describes the residual chlorine existing in water in chemical combination with ammonia or organic amines which can be found in natural or polluted waters?

- A. Combined chlorine
- B. Disinfection
- C. Free chlorine
- D. Breakpoint chlorination
- E. Total chlorine residual
- F. None of the Above

175. What term best describes the concentration of chlorine in the water after the chlorine demand has been satisfied?

- A. Chlorine Residual
- B. Disinfection
- C. Free chlorine
- D. Breakpoint chlorination
- E. Total chlorine residual
- F. None of the Above

176. What term best describes this missing term, which includes both the free and combined or chemically bound chlorine residuals?

- A. Chlorine Residual
- B. Disinfection
- C. Free chlorine
- D. Chlorine Demand
- E. Total chlorine residual
- F. None of the Above

177. What term best describes the addition of chlorine after a process or adding chlorine downstream to meet a Demand in the system?

- A. Chlorination
- B. Post-chlorination
- C. Chlorine Demand
- D. Demand
- E. Pre-chlorination
- F. None of the Above

178. Solid chlorine is about 1.5 times heavier than water and gaseous chlorine is about 2.5 times heavier than air. Atomic number of chlorine is 17. Cl is the elemental symbol and Cl₂ is the chemical formula.

- A. True
- B. False

179. What term best describes the addition of Cl₂ to the water until the Cl₂ demand is satisfied. Until all the microorganisms are killed?

- A. Organic amine(s)
- B. Disinfection
- C. Free chlorine
- D. Breakpoint chlorination
- E. Total chlorine residual
- F. None of the Above

180. _____ describes the amount of chlorine used up in a water purification system; used as a monitoring measurement by system operators.

- A. Chlorination
- B. Total
- C. Chlorine Demand
- D. Total chlorine
- E. Combined Chlorine Residual
- F. None of the Above

181. According to the text, a free chlorine residual of at least 10 mg/L should be maintained in the clear well or distribution reservoir immediately downstream from the point of post-chlorination and 2 mg/L in the distribution system to guard against backflow.

- A. True
- B. False

182. Which term best describes the total of free residual and combined residual chlorine in a water purification system; used as a monitoring measurement by system operators?

- A. Organic amine(s)
- B. Disinfection
- C. Free chlorine
- D. Breakpoint chlorination
- E. Total Chlorine Residual
- F. None of the Above

183. If chloramines are present in the municipal water supply, then total chlorine should be higher than _____.

- A. Organic amine(s)
- B. Disinfection
- C. Free chlorine
- D. Breakpoint chlorination
- E. Total chlorine residual
- F. None of the Above

184. What term best describes the method of water disinfection where gaseous, liquid, or dissolved chlorine is added to a water supply system?

- A. Chlorination
- B. Disinfection
- C. Chlorine Demand
- D. Sterilization
- E. Free chlorine
- F. None of the Above

185. What term best describes the killing of everything?

- A. Sterilization
- B. Disinfection
- C. Free chlorine
- D. Breakpoint chlorination
- E. Total chlorine meltdown
- F. None of the Above

Chlor-Alkali Membrane Process

186. The electrolysis occurs in a cell containing electrodes submerged in solutions called electrolytes. One electrode is referred to as the anode and is submerged in?

- A. Chlorination
- B. Caustic soda
- C. Chlorine ion
- D. Chlor-alkali membrane process
- E. A salt water solution
- F. None of the Above

187. The second electrode is the cathode and is submerged in a _____ solution.

- A. Oxidizing chemical(s)
- B. A salt water solution
- C. Sodium
- D. Sodium hydroxide (caustic soda)
- E. Sodium and chlorine ions
- F. None of the Above

188. Which of the following terms passes across an ion selective membrane leaving the chlorine ion to combine with a second chlorine ion, which makes a chlorine gas bubble at the anode (electrode).

- A. Chlorination
- B. Caustic soda
- C. Chlorine ion
- D. Chlor-alkali membrane process
- E. The sodium ion
- F. None of the Above

189. The membrane in the cell keeps the two solutions separate; otherwise, the chlorine gas bubble would immediately combine with the caustic soda forming?

- A. Chlorination
- B. Caustic soda
- C. Chlorine ion
- D. Chlor-alkali membrane process
- E. Sodium hypochlorite, or bleach
- F. None of the Above

Chlorine's Effectiveness

190. Chlorination depends on the chlorine demand of the water, the concentration of the chlorine solution added, the time that _____ is in contact with the organism, and water quality.

- A. Oxidizing chemical(s)
- B. Chlorine
- C. Sodium
- D. Caustic soda
- E. Sodium and chlorine ions
- F. None of the Above

191. Which term expresses a compound that is less effective in cloudy (turbid) water?

- A. Oxidizing chemical(s)
- B. Chlorination
- C. Sodium
- D. Caustic soda
- E. Sodium and chlorine ions
- F. None of the Above

192. _____ expresses is less effective as the water's pH increases (becomes more alkaline).

- A. Chlorination
- B. Caustic soda
- C. Chlorine ion
- D. Chlor-alkali membrane process
- E. Required contact time
- F. None of the Above

193. Which term is more effective as water temperature increases?

- A. Oxidizing chemical(s)
- B. Chlorination
- C. Sodium
- D. Caustic soda
- E. Sodium and chlorine ions
- F. None of the Above

Oxidation Chemistry

194. Oxidizing chemicals are often utilized in water treatment programs include: chlorine, chlorine dioxide, bromine, bromine/chlorine releasing compounds, ozone and Hydrogen peroxide.

- A. True
- B. False

195. All of the following play a role in the decision-making process: environmental and regulatory impact, _____, process contamination, and equipment capital and maintenance expense.

- A. As necessary
- B. Disinfection process
- C. System pH
- D. The primary methods used for the disinfection
- E. Economical and versatile chemicals
- F. None of the Above

196. The primary killing mechanism is oxidizing protein groups within a microorganism, these proteins are the basic components of _____ that are necessary for life-sustaining cellular processes such as respiration.

- A. Total Coliform (TC)
- B. Indicator organisms
- C. Cholera, polio, typhoid, hepatitis
- D. Cryptosporidium
- E. Essential cellular enzymes
- F. None of the Above

Chlorine Gas Section

197. When chlorine is added into the water stream, chlorine hydrolyzes into?

- A. HCL
- B. Sodium hypochlorite
- C. Bromoform
- D. Chlorine Acid
- E. Hypochlorous acid (HOCl), and hydrochloric acid (HCl)
- F. None of the Above

198. Considerably more _____ is present at a pH of 7.0 than at pH 8.5.

- A. HCl
- B. HOCl
- C. High chlorine concentrations
- D. Alkalinity
- E. Hypochlorite ion (OCl⁻)
- F. None of the Above

199. According to the text, Chlorine can be non-selective, making it very sensitive to contamination from either cooling water makeup or from in-plant process leaks. _____, organic acids and organic compounds, sulfides, iron and manganese all easily react with HOCl.
- A. Chlorine
 - B. Sodium hypochlorite
 - C. Ammonia
 - D. Chlorine gas
 - E. Hypochlorous acid (HOCl), and hydrochloric acid (HCl)
 - F. None of the Above

200. What is the term that best describes the amount of chlorine needed to react with contamination species and it must be satisfied before active HOCl is available to provide a free chlorine residual?
- A. Chlorine demand
 - B. HOCl
 - C. High chlorine concentration
 - D. Total residual
 - E. The hypochlorite ion (OCI-)
 - F. None of the Above

Pathophysiology

201. According to the text, respiratory exposure to _____ may be prolonged because its moderate water solubility may not cause upper airway symptoms for several minutes. In addition, the density of the gas is greater than that of air, causing it to remain near ground level and increasing exposure time.
- A. Hydrochloric acid
 - B. Chlorine gas
 - C. The gas
 - D. The chemical species produced
 - E. Plasma exudation
 - F. None of the Above

Mechanism of Activity

202. The mechanisms of cellular injury are believed to result from the oxidation of functional groups in cell components, from reactions with tissue water to form _____, and from the generation of free oxygen radicals.
- A. Generation of free oxygen radicals
 - B. Chlorine acid
 - C. Hydrochloric acid
 - D. A caustic effect
 - E. Hypochlorous and hydrochloric acid
 - F. None of the Above

203. Chlorine gas should be stored in vented rooms that have panic bar equipped doors.
- A. True
 - B. False

Solubility Effects

204. Which of the following terms is highly soluble in water. The predominant targets of the acid are the epithelia of the ocular conjunctivae and upper respiratory mucus membranes.
- A. Hydrochloric acid
 - B. H₂SO₄
 - C. Hypochloric acid
 - D. Sodium hypochlorite solution
 - E. Sulfuric Acid
 - F. None of the Above

Early Response to Chlorine Gas

205. If you mix ammonia with chlorine gas, this compound reacts to form?
- A. Hypochlorous acid
 - B. Chlorine gas
 - C. Hydrochloric acid
 - D. Sulfuric acid
 - E. Chloramine gas
 - F. None of the Above

Immediate Effects

206. Which of the following answers is the best choice for the immediate effects of this substance's toxicity include acute inflammation of the conjunctivae, nose, pharynx, larynx, trachea, and bronchi?

- A. Hydrochloric acid
- B. Chlorine gas
- C. Hypochlorous gas
- D. Sulfuric acid
- E. HOCL
- F. None of the Above

Pathological Findings

207. According to the text, treatment plants use _____ to reduce water levels of microorganisms that can spread disease to humans.

- A. HCl
- B. HOCl
- C. High chlorine concentrations
- D. Chlorine
- E. The hypochlorite ion (OCl⁻)
- F. None of the Above

208. Chlorine gas is the most expensive form of chlorine to use. The typical amount of chlorine gas required for water treatment is 1-16 mg/L of water. Different amounts of chlorine gas are used depending on the quality of water that needs to be treated. If the water quality is good, a higher concentration of chlorine gas will be required to disinfect the water if the contact time cannot be increased.

- A. True
- B. False

Exposure

209. After prolonged exposure, the skin can become sensitive. Sodium hypochlorite is poisonous for water organisms. It is mutagenic and very toxic when it comes in contact with Ammonium salts.

- A. True
- B. False

Routes of Exposure

Inhalation

210. Chlorine is lighter than air and may cause asphyxiation in poorly ventilated, enclosed, or high-lying areas.

- A. True
- B. False

Ingestion

211. Metabolic acidosis is rare, but has been reported following the ingestion of _____. Pulmonary complications resulting from aspiration may also be seen after ingestion.

- A. Hypochlorous Acid (HOCl)
- B. Residual disinfectant
- C. Higher levels of chlorine
- D. Sodium and calcium
- E. Household bleach
- F. None of the Above

Sources/Uses

212. According to the text, these compounds are manufactured by the chlorination of sodium hydroxide or lime.

- A. Sodium hypochlorite
- B. Chlorine gas
- C. Sodium and calcium hypochlorite
- D. Hypochlorous acid
- E. Hypochlorite solutions, powder, or concentrated vapor
- F. None of the Above

Calcium Hypochlorite Section

213. Which of the following substances comes in two forms: powder and tablets?

- A. Calcium hypochlorite
- B. Hypochlorous Acid (HOCl)
- C. Sodium hypochlorite
- D. Chlorine
- E. Hypochlorite
- F. None of the Above

214. Which substance decomposes in water to release chlorine and oxygen; sodium hypochlorite solutions can react with acids or ammonia to release chlorine or chloramine?

- A. Calcium hypochlorite
- B. Hypochlorous Acid (HOCl)
- C. Oxygen and chlorine
- D. Chlorine tablet(s)
- E. Hypochlorite ion
- F. None of the Above

Description

215. Solid chlorine stands alone as the safest form of chlorine disinfection. Requiring only minimal safety equipment for handling, users can breathe easy knowing our tablets are safe for both people and the environment. The elimination of costly scrubbers, containment, or hazard response capability, guarantees lower initial costs and reduced operating expense.

- A. True
- B. False

Accuracy

216. According to the text, this answer is an accurate dose, always yielding the stated level of available chlorine in water or very slightly over, never under.

- A. Chlorine tablet(s)
- B. Household bleach
- C. Hypochlorous Acid (HOCl)
- D. Sodium hypochlorite
- E. Calcium hypochlorite
- F. None of the Above

Effectiveness

217. Liquid Sodium hypochlorite and chlorine tablets produce Hypochlorous Acid (HOCl) and?

- A. Calcium hypochlorite
- B. Hydrochlorous Acid (HOCl)
- C. Oxygen
- D. Hypochlorite ion (OCl-) in solution
- E. Hypochlorite ion
- F. None of the Above

Safety

218. Which of the following can affect eyes, skin and mucous membranes; it is easily splashed and rots clothing?

- A. Chlorine tablet(s)
- B. Hypochlorite
- C. Chloramine
- D. Sodium dichloroisocyanurate (NaDCC)
- E. Liquid chlorine
- F. None of the Above

Corrosion

219. Which of the following are much less corrosive than liquid chlorine, which is highly corrosive to most metals?

- A. Sodium hypochlorite
- B. Hypochlorous Acid (HOCl)
- C. Oxygen and chlorine
- D. Chlorine tablet(s)
- E. Hydrochlorite
- F. None of the Above

Comparison

220. Which substance is comparable to Sodium dichloroisocyanurate (NaDCC) is their neutralization by organic matter.

- A. Chlorine tablet(s)
- B. Hypochlorite
- C. Chloramine
- D. Sodium hypochlorite (NaOCL)
- E. Hypochlorous Acid
- F. None of the Above

Health Effects

221. Which substance produces tissue injury by liquefaction necrosis. Systemic toxicity is rare, but metabolic acidosis may occur after ingestion.

- A. Ammonia
- B. Hypochlorite
- C. Chloramine
- D. Sodium dichloroisocyanurate (NaDCC)
- E. Hypochlorous Acid
- F. None of the Above

Acute Exposure

222. According to the text, the toxic effects of this compound are primarily due to the corrosive properties of the hypochlorite moiety.

- A. Calcium hypochlorite
- B. Hypochlorous Acid (HOCl)
- C. Oxygen and chlorine
- D. Sodium and calcium hypochlorite
- E. Hypochlorite ion
- F. None of the Above

Sodium Hypochlorite Solutions

223. Sodium hypochlorite solutions liberate the Toxic gases chlorine or chloramine if mixed with acid or ammonia (this can occur when bleach is mixed with another cleaning product). Thus, exposure to hypochlorite may involve exposure to these gases.

- A. True
- B. False

Potential Sequelae

224. Exposure to toxic gases generated from hypochlorite solutions can lead to reactive airways dysfunction syndrome (RADS), a chemical irritant-induced type of asthma. Chronic complications following ingestion of hypochlorite include esophageal obstruction, pyloric stenosis, squamous cell carcinoma of the esophagus, and vocal cord paralysis with consequent airway obstruction.

- A. True
- B. False

Chronic Exposure

225. Because chronic dermal can cause dermal irritation due to exposure to this substance.

- A. Chlorine tablet(s)
- B. Hypochlorite
- C. Chloramine
- D. Sodium dichloroisocyanurate (NaDCC)
- E. Hypochlorous Acid
- F. None of the Above

Chlorine-Based Disinfectants Chloramines

Chloramine Disadvantages

226. Which residual in tap water can pass through membranes in dialysis machines and directly induce oxidant damage to red blood cells?

- A. Free chlorine
- B. Chloramine
- C. Dichloramine
- D. Monochloramine
- E. Ammonia and chlorine compounds
- F. None of the Above

Chloramine Section

227. _____: $\text{NH}_3 + \text{HOCl} \rightarrow \text{NH}_2\text{Cl} + \text{H}_2\text{O}$

- A. Free chlorine
- B. Trichloramine
- C. Dichloramine
- D. Monochloramine
- E. Ammonia and chlorine compounds
- F. None of the Above

228. Free chlorine reacts with the chloramine to produce hydrogen ion, water, and _____ which will come out of solution. In the case of the monochloramine, the following reaction occurs: $2\text{NH}_2\text{Cl} + \text{HOCl} \rightarrow \text{N}_2 + 6\text{HCl} + \text{H}_2\text{O}$

- A. Free chlorine
- B. Chloramine(s)
- C. Dichloramine
- D. Nitrogen gas
- E. Ammonia and chlorine compounds
- F. None of the Above

229. Which of the following terms are formed in the pH range of 4.5 to 8.5, however, monochloramine is most common when the pH is above 8. When the pH of the water is below 4.5, the most common form of chloramine is trichloramine which produces a very foul odor.

- A. Free chlorine
- B. Chloramine(s)
- C. Dichloramine
- D. Monochloramine and dichloramine
- E. Ammonia and chlorine compounds
- F. None of the Above

Post Chlorination

230. Post chlorination is almost always done in water treatment, but can be replaced with chlorine dioxide or chloramines. In this stage, chlorine is fed to the drinking water stream which is then sent to the chlorine contact basin to allow the chlorine a long enough detention time to kill all viruses, bacteria, and protozoa that were not removed and rendered inactive in the prior stages of treatment.

- A. True
- B. False

Chlorination Equipment Requirement Section

231. Chlorine gas under pressure shall not be permitted outside the chlorine room. A chlorine room is where chlorine gas cylinders and/or ton containers are _____.

- A. Under pressure
- B. In this stage
- C. Stored
- D. At the point of solution application
- E. Dosing enough chlorine
- F. None of the Above

232. Which is the mechanical gas proportioning equipment that may or may not be located inside the chlorine room?

- A. Gas vacuum line
- B. Vacuum regulators
- C. Manual chlorine feed systems
- D. The chlorinator
- E. Injectors
- F. None of the Above

233. _____ shall be included in the gas vacuum line between the vacuum regulator(s) and the chlorinator(s) to ensure that pressurized chlorine gas does not enter the gas vacuum lines leaving the chlorine room?

- A. Gas vacuum line
- B. A gas pressure relief system
- C. Manual chlorine feed systems
- D. Mechanical gas proportioning equipment
- E. Post chlorination
- F. None of the Above

Capacity

234. Which of the following shall have the capacity to dose enough chlorine to overcome the demand and maintain the required concentration of the "free" or "combined" chlorine.

- A. The chlorinator
- B. Constant flow rate(s)
- C. Uninterrupted chlorination
- D. Automatic proportional controlled
- E. Constant pre-established dosage
- F. None of the Above

Methods of Control

235. Which of the following shall be automatic proportional controlled, automatic residual controlled, or compound loop controlled?

- A. A chlorine feed system
- B. Constant flow rate(s)
- C. Uninterrupted chlorination
- D. Automatic proportional controlled
- E. Constant pre-established dosage
- F. None of the Above

236. Which piece of chlorination equipment, the feed rate of the chlorinator is controlled by a flow proportional signal and a residual analyzer signal to maintain particular chlorine residual in the water?

- A. Gas vacuum line
- B. Compound loop control system
- C. Manual chlorine feed systems
- D. Mechanical gas proportioning equipment
- E. After post chlorination
- F. None of the Above

237. Which piece of chlorination equipment may be installed for groundwater systems with constant flow rates?

- A. Manual chlorine feed systems
- B. Constant flow rate(s)
- C. Uninterrupted chlorination
- D. Automatic proportional controlled
- E. Constant pre-established dosage
- F. None of the Above

Standby Provision

238. As a safeguard against _____, standby chlorination equipment having the capacity to replace the largest unit shall be provided.

- A. Flow change(s)
- B. Constant flow rate(s)
- C. Uninterrupted chlorination
- D. Malfunction and/or shut-down
- E. Constant pre-established dosage
- F. None of the Above

239. For uninterrupted chlorination, _____ shall be equipped with an automatic changeover system. In addition, spare parts shall be available for all chlorinators.

- A. Flow change(s)
- B. Constant flow rate(s)
- C. Gas chlorinators
- D. Automatic proportional controlled
- E. Constant pre-established dosage
- F. None of the Above

Weigh Scales

240. Scales for weighing cylinders shall be provided at all plants using chlorine gas to permit an accurate reading of total daily weight of chlorine used. At large plants, scales of the recording and indicating type are recommended. As a minimum, a platform scale shall be provided. Scales shall be of corrosion-resistant material.

- A. True
- B. False

Securing Cylinders

241. All chlorine cylinders shall be securely positioned to safeguard against movement. Tag the cylinder "empty" and store flat and chained. Ton containers may be stacked.

- A. True
- B. False

Chlorine Leak Detection

242. Which of the following related chlorine alarm equipment shall be installed at all water treatment plants using chlorine gas? Leak detection shall be provided for the chlorine rooms.

- A. Caustic soda solution reaction tanks
- B. Corrosion resistant
- C. Securely positioned
- D. Automatic chlorine leak detection
- E. Chlorine room ventilation system
- F. None of the Above

243. Which of the following related chlorine alarm equipment should be connected to a remote audible and visual alarm system and checked on a regular basis to verify proper operation?

- A. The chlorinator
- B. The facility
- C. All chlorine cylinders
- D. The chlorine gas leakage
- E. Chlorine leak detection equipment
- F. None of the Above

244. Which of the following related chlorine alarm equipment shall not automatically activate the chlorine room ventilation system in such a manner as to discharge chlorine gas?

- A. Caustic soda solution reaction tanks
- B. Corrosion resistant
- C. Leak detection equipment
- D. Automatic chlorine leak detection
- E. Chlorine room ventilation system
- F. None of the Above

245. During an emergency, if the chlorine room is occupied, the chlorine gas leakage shall be contained within the chlorine room itself in order to facilitate a proper method of clean-up.

- A. True
- B. False

246. Consideration should be given to the provision of caustic soda solution reaction tanks for absorbing the contents of leaking one-ton cylinders where such cylinders are in use.

- A. True
- B. False

247. Chlorine leak detection equipment may not be required for very small chlorine rooms with an exterior door (e.g., floor area less than 3m²).

- A. True
- B. False

248. You can use a spray solution of ammonia or a rag soaked with sulfur dioxide to detect a small Cl₂ leak. If there is a leak, the sulfur dioxide will create a white colored smoke - Sulfuric chloride.

- A. True
- B. False

Chlorine Room Design Requirements

249. Where gas chlorination is practiced, the gas cylinders and/or the ton containers up to the vacuum regulators shall be housed in a gas-tight, well illuminated, corrosion resistant and?

- A. Mechanically ventilated enclosure
- B. Corrosion resistant
- C. Securely positioned
- D. Automatic chlorine leak detection
- E. Chlorine room ventilation system
- F. None of the Above

250. The chlorinator may or may not be located inside?

- A. The chlorinator
- B. The facility
- C. All chlorine cylinders
- D. The chlorine room
- E. Chlorine leak detection equipment
- F. None of the Above

Ventilation

251. Which chlorine safety related equipment term shall have entirely separate exhaust ventilation systems capable of delivering one (1) complete air change per minute during periods of chlorine room occupancy only?

- A. Shut off
- B. The chlorine room
- C. The room
- D. Automatic chlorine leak detection
- E. Chlorine room ventilation system
- F. None of the Above

252. Which chlorine safety related equipment term should be louvered near the ceiling, the air being of such temperature as to not adversely affect the chlorination equipment.

- A. The ceiling
- B. The chlorine room
- C. Air inlets
- D. Automatic chlorine leak detection
- E. Chlorine room ventilation system
- F. None of the Above

253. Which chlorine safety related equipment term should be outside the room at all entrance or viewing points, and a clear wire-reinforced glass window?

- A. Gas chlorine room
- B. The chlorine room
- C. Chlorine room ventilation system
- D. Automatic chlorine leak detection
- E. Separate switches for fans and lights
- F. None of the Above

Heating

254. Chlorine rooms shall have _____, if a forced air system is used to heat the building.

- A. Gas chlorine room
- B. Separate heating systems
- C. The room
- D. Automatic chlorine leak detection
- E. Chlorine room ventilation system
- F. None of the Above

255. Which chlorine safety related equipment term shall be protected to ensure that the chlorine maintains its gaseous state when entering the chlorinator.

- A. Cylinders or containers
- B. Corrosion resistant
- C. Securely positioned
- D. Automatic chlorine leak detection
- E. Chlorine room ventilation system
- F. None of the Above

Storage of Chlorine Cylinders

256. Which of the following terms may be provided to simply store the chlorine gas cylinders, with no connection to the line?

- A. Cylinders or containers
- B. The outside of the room
- C. A separate storage room
- D. Uncontrolled release of spilled gas
- E. Air inlets
- F. None of the Above

257. Which chlorine safety related equipment term shall have provision for ventilation at thirty air changes per hour?

- A. A panic button
- B. The chlorine room
- C. Scrubber(s)
- D. The chlorine gas storage room
- E. The chlorine cylinder storage room
- F. None of the Above

258. Sometimes entry in very large facilities, may be through a vestibule from outside in to?

- A. Cylinders or containers access
- B. The outside of the room
- C. Chlorine rooms
- D. Uncontrolled release of spilled gas
- E. Air inlets
- F. None of the Above

Scrubbers

259. According to the text, facilities located within residential or densely populated areas, consideration shall be given to provide scrubbers for?

- A. A panic button
- B. The chlorine room
- C. Scrubber(s)
- D. The chlorine gas storage room
- E. The chlorine cylinder storage room
- F. None of the Above

260. Which term must be satisfied before chlorine becomes available to accomplish disinfection?

- A. Combined residual
- B. Free chlorine residual
- C. Demand for chlorine
- D. Total chlorine
- E. Free chlorine
- F. None of the Above

261. Which term means the amount of chlorine required to produce a residual of 0.1 mg/l after a contact time of fifteen minutes as measured by Iodometric method of a sample at a temperature of twenty degrees in conformance with Standard methods?

- A. Combined residual
- B. Free chlorine residual
- C. Chlorine Demand
- D. Total chlorine
- E. Break point chlorination
- F. None of the Above

Chlorine Health Hazard Section

262. Which term expresses low levels of chlorine results in eye, nose, and throat irritation, sneezing, Excessive salivation, general excitement, and restlessness?

- A. Rambling
- B. Inhalation
- C. Acute exposure
- D. Chronic exposure
- E. Immediate attention after inhalation
- F. None of the Above

263. _____ expresses low levels of chlorine gas can result in a dermatitis known as chloracne, tooth enamel corrosion, coughing, sore throat, hemoptysis and increased susceptibility to tuberculosis.

- A. Rambling
- B. Inhalation
- C. Acute exposure
- D. Chronic exposure
- E. Immediate attention after inhalation
- F. None of the Above

Inhalation

264. The nose and throat may become irritated; a stinging and burning sensation may be experienced. Immediate fatalities can occur as a result of suffocation. Delayed fatalities can occur as a result of pulmonary edema (fluid in the lungs). For this reason, rest and immediate attention after inhalation is important.

- A. True
- B. False

265. If breathing has stopped, give artificial respiration; if breathing is difficult, give oxygen if equipment and trained personnel are available. If exposed person is breathing, place in a comfortable position and keep person warm and at rest until medical assistance becomes available.

- A. True
- B. False

Eye/Skin Contact

266. Liquid and concentrated gas could produce severe?

- A. Burns and injury on contact
- B. Plasma exudation
- C. General excitement
- D. Chronic exposure to low levels of chlorine gas
- E. Inhalation due to stress
- F. None of the Above

Eye

267. If you get chlorine in the eye, pour a gentle stream of _____ through the affected eye for at least 15 minutes. Contact the poison control center, emergency room or physician right away as further treatment will be necessary.

- A. Liquid
- B. Warm water
- C. Milk
- D. Salt water
- E. Cold water
- F. None of the Above

Skin

268. If you get chlorine on the skin, run _____ over the affected area for 15 minutes.

- A. A gentle stream of water
- B. Warm water
- C. Milk
- D. Salt water
- E. Cold water
- F. None of the Above

Chronic

269. Repeated exposures to chlorine gas can result in a loss of ability to detect the odor of chlorine. Long term exposures may cause damage to teeth and inflammation or?

- A. Chlorine gas toxicity
- B. Plasma exudation
- C. Pulmonary edema
- D. Ulceration of the nasal passages
- E. Noncardiogenic pulmonary edema
- F. None of the Above

Pre-hospital Management

270. Rescue personnel are at low risk of noncardiogenic pulmonary edema contamination from victims who have been exposed only to gases released from hypochlorite solutions. However, clothing or skin soaked with industrial-strength bleach or similar solutions may be corrosive to rescuers and may release harmful gases.

- A. True
- B. False

271. Ingestion of hydrochlorite solutions rarely causes pain in the mouth or throat, dysphagia, stridor, drooling, odynophagia, and vomiting.

- A. True
- B. False

272. Chronic exposure to gases released from ammonia solutions can cause coughing, eye and nose irritation, lacrimation, and a burning sensation in the chest.

- A. True
- B. False

Hot Zone

273. Which term is the area that rescuers should be trained and appropriately attired before entering?

- A. Support Zone
- B. Warm zone
- C. Chemical-protective clothing area
- D. Decontamination area
- E. Hot Zone
- F. None of the Above

Rescuer Protection

274. Which of the following terms is irritating to the skin and eyes and in some cases may release toxic gases?

- A. Hydrothromine
- B. Hypochlorite
- C. Chloramine
- D. Sodium dichloroisocyanurate (NaDCC)
- E. Ammonia
- F. None of the Above

275. Positive-pressure, self-contained breathing apparatus (SCBA) is recommended in response to situations that involve exposure to potentially unsafe levels of?

- A. Chlorine tablet(s)
- B. Hypochlorite
- C. Chlorine gas
- D. Solid hypochlorite or concentrated solutions
- E. Hypochlorous Acid
- F. None of the Above

276. Chemical-protective clothing should be worn due to the risk of skin irritation and burns from direct contact with _____.

- A. Chlorine tablet(s)
- B. Hypochlorite
- C. Chlorine gas
- D. Solid hypochlorite or concentrated solutions
- E. Hypochlorous Acid
- F. None of the Above

ABC Reminders

277. If a person is over taken with chlorine gas exposure, quickly establish a _____, ensure adequate respiration and pulse.

- A. Support Zone
- B. Patient airway
- C. Chemical-protective clothing
- D. Delay decontamination
- E. Hot Zone to the Decontamination Zone
- F. None of the Above

Victim Removal

278. During the chlorine evacuation, if victims can walk, lead them out of the?

- A. Decontamination area
- B. Hot Zone
- C. Chemical-free zone
- D. Chemically contaminated zone
- E. Hot Zone to the Decontamination Zone
- F. None of the Above

Decontamination Zone

279. Victims may be transferred immediately to the?

- A. Support Zone
- B. Patient Zone
- C. Chemical free zone
- D. Decontamination area
- E. Hot Zone to the Decontamination Zone
- F. None of the Above

Rescuer Protection

280. Decontamination may be conducted by personnel wearing a lower level of protection than that worn in the _____, if exposure levels are determined to be safe.

- A. Support Zone
- B. Patient Zone
- C. Chemical free zone
- D. Decontamination area
- E. Hot Zone
- F. None of the Above

ABC Reminders

281. Quickly establish a _____, ensure adequate respiration and pulse.

- A. Support Zone
- B. Patient airway
- C. Hot Zone
- D. Decontamination zone
- E. Chemical-protective clothing dressing area
- F. None of the Above

Basic Decontamination

282. During a chlorine leak, _____ is critical.

- A. Decontamination
- B. Hot Zone
- C. Chemical-protective clothing
- D. Rapid decontamination
- E. Hot Zone to the Decontamination Zone
- F. None of the Above

In Cases of Ingestion, Do Not Induce Emesis or Offer Activated Charcoal.

283. During a chlorine leak, victims who are conscious and able to swallow should be given 4 to 8 ounces of?

- A. Liquid
- B. Warm water
- C. Milk only
- D. Water or milk
- E. Cold water
- F. None of the Above

284. During a chlorine leak, consider appropriate _____ of chemically contaminated children at the exposure site. Provide reassurance to the child during decontamination, especially if separation from a parent occurs.

- A. Decontamination
- B. Hot Zone
- C. Chemical-protective clothing
- D. Management
- E. Hot Zone to the Decontamination Zone
- F. None of the Above

Chlorine Dioxide Section

285. ClO₂ generation uses _____ and chlorine gas.

- A. Chlorine dioxide (ClO₂)
- B. Sodium chlorite (NaClO₂)
- C. Hypochlorous acid
- D. NaOCl and HCl in place of chlorine gas
- E. Ozone
- F. None of the Above

286. Chlorine gas is educted into a motive water stream in a ClO₂ generator forming?

- A. Hypochlorous acid
- B. HOCl and HCl
- C. Chlorine dioxide
- D. Sodium chlorate (NaClO₃) and sulfuric acid
- E. Sodium thiosulfate
- F. None of the Above

287. This compound is pumped into the stream and allowed to react in a generating column to produce ClO₂?

- A. Hypochlorous acid
- B. HOCl and HCl
- C. Chlorine dioxide
- D. Sodium chlorite
- E. Sodium thiosulfate
- F. None of the Above

288. _____ does not hydrolyze in water as chlorine does and with it. It remains fully active in a pH range far broader than chlorine or sodium hypochlorite.

- A. Sodium chlorite (NaClO₂)
- B. Chlorine gas
- C. Chlorine dioxide or ClO₂
- D. Sodium chlorate (NaClO₃)
- E. NaOCl and HCl
- F. None of the Above

289. Which of the following compound(s) remains a gas in water, it does not have the corrosive tendencies of chlorine gas?

- A. Sodium chlorite (NaClO₂)
- B. Chlorine gas
- C. Chlorine dioxide or ClO₂
- D. Sodium chlorate (NaClO₃)
- E. NaOCl and HCl
- F. None of the Above

290. Which of the following compound(s) is a dissolved gas in water; there is no mineral acid or caustic soda formation as happens when using HOCl.

- A. ClO_2
- B. Sodium chlorite (NaClO_2)
- C. Hypochlorous acid
- D. NaOCl and HCl in place of chlorine gas
- E. Heavily pH-dependent
- F. None of the Above

291. Which of the following compound(s) tends to be much less, if not totally non-reactive, with many organic and inorganic compounds.

- A. ClO_2
- B. Sodium chlorite (NaClO_2)
- C. Hypochlorous acid
- D. NaOCl and HCl in place of chlorine gas
- E. Heavily pH-dependent
- F. None of the Above

292. Which of the following compound(s) is much less aggressive to traditional corrosion inhibitors?

- A. Sodium chlorite (NaClO_2)
- B. Chlorine gas
- C. Chlorine dioxide or ClO_2
- D. Sodium chlorate (NaClO_3)
- E. NaOCl and HCl
- F. None of the Above

293. _____ is heavily pH-dependent, because as system pH increases, there is a correspondingly rapid decrease in the concentration of the biocidally active species.

- A. Chlorine dioxide (ClO_2)
- B. Sodium chlorite (NaClO_2)
- C. Hypochlorous acid
- D. NaOCl and HCl in place of chlorine gas
- E. Sodium chlorate (NaClO_3) and sulfuric acid
- F. None of the Above

294. Which of the following compound(s) is a non-specific oxidant which readily reacts with various organic and inorganic compounds that may be present in a cooling water system.

- A. Chlorine dioxide (ClO_2)
- B. Sodium chlorite (NaClO_2)
- C. Hypochlorous acid
- D. NaOCl and HCl in place of chlorine gas
- E. Sodium chlorate (NaClO_3) and sulfuric acid
- F. None of the Above

295. _____ is considerably more selective than chlorine in the presence of various compounds, which allows it to be more effective in contaminated systems?

- A. ClO_2
- B. Sodium chlorite (NaClO_2)
- C. Hypochlorous acid
- D. NaOCl and HCl in place of chlorine gas
- E. Sodium chlorate (NaClO_3) and sulfuric acid
- F. None of the Above

296. Which of the following compound(s), can be in fact, be two-and-one-half times more reactive than chlorine?

- A. ClO_2
- B. Sodium chlorite (NaClO_2)
- C. Hypochlorous acid
- D. NaOCl and HCl in place of chlorine gas
- E. Sodium chlorate (NaClO_3) and sulfuric acid
- F. None of the Above

297. Which of the following terms as a water disinfectant increased in the 1970s when it was discovered that it did not promote THM formation?

- A. Sulfur Dioxide
- B. Chlorine gas
- C. Chlorine dioxide
- D. Sodium chlorate (NaClO_3) and sulfuric acid
- E. UV
- F. None of the Above

298. _____ was used in the paper industry, has been an acceptable and effective alternative to chlorination in cooling systems?

- A. Chlorine dioxide (ClO_2)
- B. Sodium chlorite (NaClO_2)
- C. Hypochlorous acid
- D. NaOCl and HCl in place of chlorine gas
- E. Sodium thiosulfate
- F. None of the Above

299. This compound is a yellow-green gas with an irritating odor not unlike chlorine.

- A. Sodium thiosulfate
- B. Chlorine
- C. Chlorine dioxide
- D. Sodium chlorate (NaClO_3) and sulfuric acid
- E. Ozone
- F. None of the Above

300. This compound cannot be compressed and shipped in a container, so it must be generated on site.

- A. Sodium thiosulfate
- B. Chlorine
- C. Chlorine dioxide
- D. Sodium chlorate (NaClO_3) and sulfuric acid
- E. Ozone
- F. None of the Above

301. Which of the following compound(s) under efficient generation, THMs are not formed and THM precursor(s) are reduced. In one application, THM formation was reduced from 34 mg/l to 1 mg/l?

- A. ClO_2
- B. Sodium chlorite (NaClO_2)
- C. Hypochlorous acid
- D. NaOCl and HCl in place of chlorine gas
- E. Sodium chlorate (NaClO_3) and sulfuric acid
- F. None of the Above

302. Which of the following compound(s) is formed from the dissolution of chlorine gas or sodium hypochlorite in water, has satisfactorily controlled microorganisms in cooling water systems?

- A. Chlorine tablet(s)
- B. Hydrochlorous acid
- C. Chlorine gas
- D. Solid hypochlorite or concentrated solutions
- E. Hypochlorous Acid
- F. None of the Above

303. The effects of _____ on hypochlorous acid and its reactivity with a variety of compounds both combine to vastly diminish its effectiveness in contaminated, high-pH cooling water systems. Conversely, chlorine dioxide remains completely pH-independent in the range where recirculating and once-through cooling systems are typically operated.

- A. THM precursor(s)
- B. Chlorine gas
- C. Chlorine dioxide
- D. pH
- E. THM formation
- F. None of the Above

Water Disinfection Methods Review

304. Water systems add _____ to destroy microorganisms that can cause disease in humans.

- A. Alkalinity and pH
- B. Hydrogen peroxide
- C. Hypochlorous acid
- D. Oxidizing and biocidal properties
- E. Disinfectants
- F. None of the Above

305. The following primary methods of disinfection are chlorination, chloramines, ozone, and ultraviolet light. Other disinfection methods include chlorine dioxide, _____.

- A. Alkalinity and pH
- B. Hydrogen peroxide
- C. Hypochlorous acid
- D. Oxidizing and biocidal properties
- E. Potassium permanganate, and nanofiltration
- F. None of the Above

306. Since certain forms of chlorine react with _____ naturally present in many water sources to form harmful chemical by-products.

- A. Alkalinity and pH
- B. Organic material
- C. Hypochlorous acid
- D. Oxidizing and biocidal properties
- E. Hazardous trihalomethanes (THM)
- F. None of the Above

Physical Methods

307. Formation of _____ in water and wastewater effluent treated with chlorine has prompted research to seek alternative disinfecting methods that would minimize environmental and public health impacts.

- A. Alkalinity
- B. Mutagenic and carcinogenic agents
- C. Hypochlorous acid
- D. Oxidizing and biocidal properties
- E. Hazardous trihalomethanes (THM)
- F. None of the Above

Chemical Methods

308. Chemical methods depend mostly on selected chemicals with oxidizing and biocidal properties. Their practical applications range from removing _____ to disinfecting water supplies, wastewater treatment effluent, or industrial waters.

- A. Alkalinity and pH
- B. Undesirable constituents
- C. Hypochlorous acid
- D. Oxidizing and biocidal properties
- E. Hazardous trihalomethanes (THM)
- F. None of the Above

309. Which of the following compound(s) used for disinfection, other than chlorine and some of its compounds, potassium permanganate, and hydrogen peroxide?

- A. Ammonia
- B. Sodium chlorite (NaClO_2)
- C. Hydrochlorous acid
- D. NaOCl and HCl in place of chlorine gas
- E. Ozone
- F. None of the Above

310. Ozonation enhances the _____ despite its inherent weakness in leaving practically no residual in the distribution system.

- A. Effectiveness and cost
- B. Protecting public health
- C. Mode of disinfection
- D. Coagulation process
- E. Superiority over chlorination
- F. None of the Above

Chlorination and Dechlorination

311. _____ and some of its derivatives will continue as an integral part of the disinfection process in water and wastewater treatment.

- A. Chlorine tablet(s)
- B. Hydrochlorous acid
- C. Chlorine
- D. Solid hypochlorite or concentrated solutions
- E. Hypochlorous Acid
- F. None of the Above

pH Scale

312. Alkalinity is the capacity of water to increase acids. This increase is caused by the water's content of carbonate, bicarbonate, hydroxide and occasionally borate, silicate and phosphate.

- A. True
- B. False

313. pH is an expression of the intensity of the basic or acid condition of a liquid. EPA has a suggested range of 5.5 to 7.5 for pH (called a primary maximum contaminant level or MCL).

- A. True
- B. False

Alternative Disinfectants Chapter 4

Ultraviolet Disinfection

314. The microorganisms spend maximum time and contact with the outside of the quartz tube and the source of the?

- A. Sterilizer
- B. UV rays
- C. UV disinfection
- D. UV reactor
- E. Electromagnetic energy
- F. None of the Above

315. According to the text, the _____ will consist of a various number of lamps and tubes, depending upon the quantity of water to be treated.

- A. UV sterilizer
- B. UV rays
- C. UV disinfection
- D. UV reactor
- E. Electromagnetic energy
- F. None of the Above

316. Ensuring that the _____ maintains good contact with the water requires control of the water level within the channel to ensure that the UV is making total contact at the designed depths.

- A. UV
- B. Contact
- C. Channel
- D. UV reactor
- E. Ballasts and shields
- F. None of the Above

317. Heat is generated by the electric components of the UV system, adequate ventilation and cooling must be applied to the _____ to reduce heat build-up, otherwise the ballasts will fail.

- A. UV arrays
- B. UV rays
- C. UV disinfection
- D. UV reactor
- E. Electromagnetic energy
- F. None of the Above

318. Because of the great electrical consumption of this system, combined with the cost of routine replacement of _____, should be considered against other systems.

- A. UV capacitor
- B. UV Flux
- C. UV disinfection
- D. UV reactor
- E. Ballasts and shields
- F. None of the Above

319. Disadvantages include the effects of turbidity in the water reducing the infiltration and therefore the effectiveness of ballasts and shields and the need to provide an effective cleaning and replacement program for the UV components.

- A. True
- B. False

320. The effective use of Ultraviolet treatment, the water to be disinfected can contain suspended solids. The water does not need to be colorless and can contain colloids, iron, manganese, taste, and odor.

- A. True
- B. False

Strongest Oxidizing Agent

321. This compound is obtained by passing a flow of air or oxygen between two electrodes that are subjected to an alternating current in the order of 10,000 to 20,000 volts.

- A. Chloramine
- B. Liquid Ozone
- C. Ozone
- D. Oxygen and nascent oxygen
- E. O₂
- F. None of the Above

322. This compound is a light blue gas at room temperature.

- A. Chloramine
- B. Liquid Ozone
- C. Ozone
- D. Oxygen and nascent oxygen
- E. O₂
- F. None of the Above

323. Ozone has a _____ similar to that sometimes noticed during and after heavy electrical storms. In use, ozone breaks down into oxygen and nascent oxygen.

- A. Self-policing pungent odor
- B. THMs
- C. Light blue gas
- D. Oxygen and nascent oxygen
- E. Strongest oxidizing agent
- F. None of the Above

324. Ozone does not form chloramines or _____, and while it may destroy some THMs, it may produce others when followed by chlorination.

- A. Carcinogens
- B. THMs
- C. Complete disinfectant
- D. Oxygen and nascent oxygen
- E. Flocculation and coagulation
- F. None of the Above

325. Ozone falls into the same category as other disinfectants in that it can produce?

- A. Carcinogens
- B. THMs
- C. DBPs
- D. Oxygen and nascent oxygen
- E. Strongest oxidizing agent
- F. None of the Above

326. This compound is very unstable and can readily explode. As a result, it is not shipped and must be manufactured on-site.

- A. Chloramine
- B. Liquid Ozone
- C. Ozone
- D. Oxygen and nascent oxygen
- E. O₂
- F. None of the Above

327. Each water has its own?

- A. Carcinogens
- B. THMs
- C. Ozone demand
- D. Oxygen and nascent oxygen
- E. Strongest oxidizing agent
- F. None of the Above

Alternate Disinfectants Section Summary

Chloramines

328. This compound is a very weak disinfectant for Giardia and virus reduction. It is recommended that it be used in conjunction with a stronger disinfectant. It is best utilized as a stable distribution system disinfectant.

- A. Chlorine
- B. Chloramine
- C. Ozone
- D. Oxygen and nascent oxygen
- E. Strongest oxidizing agent
- F. None of the Above

329. In the production of chloramines, the ammonia residuals in the finished water, when fed in excess of stoichiometric amount needed, should be limited to inhibit growth of?

- A. Cryptosporidium
- B. Chlorine-based disinfectants
- C. Giardia lamblia
- D. An emerging parasitic protozoan pathogen
- E. Nitrifying bacteria
- F. None of the Above

Chlorine Dioxide

330. According to the text, Chlorine dioxide provides good _____ protection but its use is limited by the restriction on the maximum residual of 0.5 mg/L ClO_2 /chlorite/chlorate allowed in finished water.

- A. Pre-disinfectant
- B. ClO_2 /chlorite/chlorate
- C. Level of
- D. Chlorine residual
- E. Giardia and virus
- F. None of the Above

331. Where chlorine dioxide is approved for use as an oxidant, the preferred method of generation is to entrain chlorine gas into a packed reaction chamber with a?

- A. Pre-disinfectant
- B. ClO_2 /chlorite/chlorate
- C. An oxidant
- D. Total residual oxidants
- E. 25% aqueous solution of sodium chlorite (NaClO_2)
- F. None of the Above

332. Because dry sodium chlorite is dangerous and can be _____ in feed equipment if leaking solutions or spills are allowed to dry out.

- A. Prone to fire
- B. Choking risk
- C. An oxidant
- D. Oxidant
- E. Explosive and can cause fires
- F. None of the Above

Ozone

333. Which term must be determined for the ozone basin alone; an accurate T10 value must be obtained for the contact chamber, residual levels measured through the chamber and an average ozone residual calculated?

- A. Ozone CT (Contact time)
- B. Free and/or combined chlorine
- C. Residual levels
- D. Contact time
- E. Strongest oxidizing agent
- F. None of the Above

334. Ozone may also be used as _____ for removal of taste and odor, or may be applied as a pre-disinfectant.

- A. An oxidant
- B. Free and/or combined chlorine
- C. Residual levels
- D. System residual
- E. Strongest oxidizing agent
- F. None of the Above

Waterborne Pathogens Chapter 5

How Diseases are Transmitted.

335. Waterborne pathogens are primarily spread by the?

- A. Fecal-oral, or feces-to-mouth, route
- B. Dermal to fecal route
- C. Oral to fecal route
- D. Influenza route
- E. Waterborne mishaps
- F. None of the Above

336. When infected humans or animals pass the bacteria, viruses, and _____ in their stool, pathogens may get into water and spread disease.

- A. Fecal Coliform and E coli
- B. Protozoa
- C. Macroorganisms
- D. Cryptosporidiosis
- E. Bioslime
- F. None of the Above

337. For another person to become infected, he or she must take that pathogen in through the mouth.

- A. True
- B. False

338. This term means when in nature it is different from other types of pathogens such as the viruses that cause influenza (the flu) or the bacteria that cause tuberculosis.

- A. Fecal Coliform and E coli
- B. Giardia lamblia
- C. Microorganism(s)
- D. Waterborne Pathogen(s)
- E. Coliform bacteria
- F. None of the Above

339. According to the text, _____ are spread by secretions that are coughed or sneezed into the air by an infected person.

- A. Fecal Coliform and E coli
- B. Giardia lamblia
- C. Microorganisms
- D. Influenza virus and tuberculosis bacteria
- E. Coliform bacteria
- F. None of the Above

Maximum Contaminant Levels (MCLs)

340. State and federal laws establish standards for drinking water quality. Under normal circumstances when these guidelines are being met, the water is somewhat safe to drink with little threat to human health.

- A. True
- B. False

341. EPA had developed standards that are known as maximum contaminant levels (MCL). When a particular contaminant exceeds this term a potential health threat may occur.

- A. Coliform bacteria count
- B. MCL
- C. Standards
- D. HPC
- E. CFU
- F. None of the Above

342. This acronym generally expresses properties of the contaminants, risk assessments and factors, short-term (acute) exposure and long-term (chronic) exposure.

- A. Coliform bacteria
- B. MCLs
- C. Standards
- D. HPC
- E. CFU
- F. None of the Above

343. When you as the operator take samples to ensure your water is in compliance with the MCL, there are two types of _____ for coliform bacteria.

- A. Coliform bacteria
- B. MCLs
- C. Standards
- D. MCL violations
- E. CFU
- F. None of the Above

344. The first type of _____ is for total coliform; the second is an acute risk to health violation characterized by the confirmed presence of fecal coliform or E. coli.

- A. Coliform bacteria
- B. MCLs
- C. Standards
- D. MCL violations
- E. CFU
- F. None of the Above

Heterotrophic Plate Count HPC

345. Heterotrophic Plate Count (HPC) --- formerly known as the Standard plate count, is a procedure for estimating the number of live heterotrophic bacteria and measuring changes during water treatment and distribution in water or in swimming pools.

- A. True
- B. False

Spread Plate Method

346. During the Spread Plate Method, colonies can be transferred quickly, and _____ easily can be discerned and compared to published descriptions.

- A. Colonies grow
- B. Surface growth
- C. Low counts
- D. Heterotrophic organisms will grow
- E. Colony morphology
- F. None of the Above

Membrane Filter Method

347. This method permits testing large volumes of _____ and is the method of choice for low-count waters.

- A. Colonies
- B. Surface water
- C. Low-turbidity water
- D. Heterotrophic organisms
- E. MCL
- F. None of the Above

Heterotrophic Plate Count (Spread Plate Method)

348. After an incubation period, a bacteriological colony count provides an estimate of the concentration of heterotrophs in the sample of interest. The R2A agar provides a medium that will support a large variety of?

- A. Colonies
- B. Bugs
- C. Germs
- D. Heterotrophic bacteria
- E. MCL
- F. None of the Above

Total Coliforms

349. This MCL is based on the presence of total coliforms, and compliance is on a daily or weekly basis, depending on your water system type and state rule.

- A. True
- B. False

350. For systems that collect fewer than _____ samples per month, no more than one sample per month may be positive. In other words, the second positive result (repeat or routine) in a month or quarter results in a MCL violation.

- A. 5
- B. 10
- C. 100
- D. 200
- E. 40
- F. None of the Above

351. For systems that collect _____ or more samples per month, no more than five (5) percent may be Positive, check with your state drinking water section or health department for further instructions.

- A. 5
- B. 10
- C. 100
- D. 200
- E. 40
- F. None of the Above

Acute Risk to Health (Fecal coliforms and E. coli)

352. A routine analysis shows total and _____ is followed by a repeat analysis which indicates total coliform present.

- A. Routine analysis
- B. Drinking water violation
- C. MCL violation
- D. Human health violations
- E. Fecal coliform or E. coli present
- F. None of the Above

353. Which of the following terms requires the water system to provide public notice via radio and television stations in the area?

- A. Routine analysis violation
- B. Drinking water rule violation
- C. MCL violation
- D. Human health violation
- E. Acute health risk violation
- F. None of the Above

354. According to the text, the type of contamination can pose an immediate threat to human health and notice must be given as soon as possible, but no later than 24 hours after notification from your laboratory of the test results.

- A. True
- B. False

Public Notice

355. There shall be certain information, be issued properly and in a timely manner, and contain certain _____ on the public notice.

- A. Legal analysis
- B. Drinking water rule information
- C. NOVs
- D. Mandatory language
- E. Fecal language
- F. None of the Above

356. If there is a(n) _____ present to users, the timing and place of posting of the public notice may have different priorities.

- A. Routine analysis
- B. Drinking water rule
- C. Acute risk
- D. Human health violation
- E. Fecal coliform or E. coli present
- F. None of the Above

The following are acute violations:

357. Which is violation of nitrate?

- A. Presence
- B. MCL
- C. MCLG
- D. Count
- E. Acute violations
- F. None of the Above

358. Concerning total coliforms - when fecal coliforms or E. coli are present in the distribution system and is a violation of the?

- A. Presence
- B. MCL
- C. MCLG
- D. Count
- E. Acute violations
- F. None of the Above

359. Any outbreak of _____, as defined by the rules.

- A. Total coliforms
- B. MCL
- C. Waterborne disease
- D. Radioactive bacteria
- E. Acute violations
- F. None of the Above

E-Coli Section

360. Escherichia coli. There are several pathogenic strains of Escherichia coli, which are classified under enterovirulent E. coli. They are enterohemorrhagic, enteroinvasive, enterotoxigenic, enteropathogenic, and enteroaggregative.

- A. True
- B. False

361. Prevention strategies for E. coli O157:H7 include _____, halogenation of water, or boiling water for one minute.

- A. Primary protection
- B. Source protection
- C. Sodium chlorite
- D. Eliminating snails with a molluscicide
- E. Backflow prevention
- F. None of the Above

362. What is the bacterial disease caused by the Salmonella species that causes diarrheal illness?

- A. Beaver fever
- B. Escherichia coli O157:H7
- C. Bacteria
- D. Pseudomonas
- E. Salmonellosis
- F. None of the Above

Repeat Sampling

363. Repeat sampling replaces the old check sampling with a more comprehensive procedure to try to _____ areas in the system.

- A. Double check the routine sample
- B. Identify problem
- C. Originate the sampling location
- D. Sample
- E. Calculate MCL compliance
- F. None of the Above

364. According to the text, whenever a Routine sample is total coliform or fecal coliform present, a set of repeat samples must be collected within how many hours after being notified by the laboratory.

- A. 12
- B. 24
- C. 48
- D. 10
- E. 2
- F. None of the Above

365. The follow-up for repeat sampling is: If only one _____ per month or quarter is required, four (4) repeat samples must be collected.

- A. Routine sample
- B. Surface water sample
- C. Original sample
- D. Sample
- E. MCL sample
- F. None of the Above

366. For systems collecting two (2) or more routine samples per month, three (3) _____ must be collected.

- A. Routine samples
- B. Surface water samples
- C. Samplers
- D. Repeat samples
- E. MCL compliance calculations
- F. None of the Above

367. Repeat samples must be collected from: The original sampling location of the?

- A. Routine sample
- B. Surface water
- C. Coliform present sample
- D. Sample
- E. MCL area
- F. None of the Above

368. Samples should be taken elsewhere in the _____ or at the wellhead, if necessary.

- A. Sewage system
- B. Surface system
- C. Sampling location
- D. Distribution system
- E. MCL compliance calculation
- F. None of the Above

369. In a very small system if the system has only _____, the repeat samples must be collected from the same sampling location over a four-day period or on the same day.

- A. Routine water
- B. Surface water
- C. One sampling location
- D. One service connection
- E. MCL compliance zone
- F. None of the Above

370. If a repeat sample is necessary, all repeat samples are included in the?

- A. Routine sample
- B. Surface water
- C. Original sampling location
- D. Sample
- E. MCL compliance calculation
- F. None of the Above

Positive or Coliform Present Results

371. According to the text, if you are notified of a positive test result you need to contact either the Drinking Water Program or your local county health department within 24 hours, or by the next business day after the?

- A. Results are reported to you
- B. Positive violation
- C. Repeat sampling immediately
- D. Sample violation
- E. MCL compliance violation
- F. None of the Above

372. Which of the following is very important to initiate as the corrective measures will be based on those results?

- A. Storage and distribution
- B. Repeat sampling immediately
- C. Upgrading of the wellhead area
- D. Perform routine procedures
- E. Corrective measures
- F. None of the Above

Legionella

373. Legionnaire's disease, which causes a severe pneumonia, and the second, _____, which is a non-pneumonia illness; it's typically an influenza-like illness, and it's less severe.

- A. Campylobacter
- B. Pathogen
- C. Pontiac fever
- D. Typhoid fever
- E. Shigella dysenteriae
- F. None of the Above

374. Which pathogen is naturally found in water, both natural and artificial water sources?

- A. Campylobacter
- B. Legionella
- C. Pontiac fever
- D. Typhoid fever
- E. Hydrodysenteriae
- F. None of the Above

375. Legionella, prevention. Legionella in water systems. Hot water in tanks should be maintained between _____ degrees Centigrade.

- A. 81 to 100
- B. 110 to 210
- C. 75 – 212
- D. 71 and 77
- E. 75 and 85
- F. None of the Above

Pseudomonas

376. Pseudomonas, the basics. It's a protozoon. It is caused by visual contact with water. It can cause dermatitis, which is an inflammation of the skin, or it can cause otitis, which is an infection of the ear.

- A. True B. False

Norovirus

377. Humans are the reservoir for the Norovirus. Prevention strategies for this pathogen include?

- A. Maintaining water systems D. Containment protection
B. Source protection E. Internal protection
C. Chlorine monoxide F. None of the Above

Cryptosporidium

378. Cryptosporidium causes diarrheal illness known as?

- A. Vomiting D. Cryptosporidiosis
B. Hemorrhagic colitis E. Salmonellosis
C. Diarrhea F. None of the Above

379. Cryptosporidium is typically associated with animals and humans, and it can be acquired through consuming fecally contaminated food, contact with fecally contaminated soil and water.

- A. True B. False

Giardia

380. Giardia prevention strategies for this pathogen include _____; filtration, coagulation, and halogenation of drinking water.

- A. Maintaining hot water systems D. Primary protection
B. Source protection E. Secondary measurements
C. Sulfur dioxide F. None of the Above

Schistosomatidae

381. Schistosomatidae, the basics. It is a parasite. It is acquired through dermal contact, cercarial dermatitis. It is commonly known as?

- A. Swimmer's itch D. Pseudomonas
B. Beaver fever E. Salmonellosis
C. Hemorrhagic colitis F. None of the Above

Respiratory Protection Section Chapter 5

382. This definition means a respiratory device that is designed for use only during escape from hazardous atmospheres.

- A. Filter or Air-Purifying Element D. End-Of-Service-Life Indicator
B. Escape Gas Mask E. Filtering Facepiece
C. Escape Only Respirator F. None of the Above

383. _____ means a system that warns the respirator user of the approach of the end of adequate respiratory protection; for example, that the sorbent is approaching saturation or is no longer effective.

- A. Filter or Air-Purifying Element D. End-Of-Service-Life Indicator
B. Escape Gas Mask E. Filtering Facepiece
C. Escape Only Respirator F. None of the Above

384. This definition means a gas mask that consists of a half-mask facepiece or mouthpiece, a canister, and associated connections, and that is designed for use during escape-only from hazardous atmospheres.

- A. Filter or Air-Purifying Element
- B. Escape Gas Mask
- C. Escape Only Respirator
- D. End-Of-Service-Life Indicator
- E. Filtering Facepiece
- F. None of the Above

385. _____ means a solid, mechanically produced particle with a size ranging from submicroscopic to macroscopic.

- A. Employee Exposure
- B. Atmosphere-Supplying Respirator
- C. Emergency Respirator Use Situation
- D. Emergency situation
- E. Immediately Dangerous to Life or Health
- F. None of the Above

386. This definition means a respirator intended to be used only for emergency exit.

- A. Filter or Air-Purifying Element
- B. Escape Gas Mask
- C. Escape Only Respirator
- D. End-Of-Service-Life Indicator
- E. Filtering Facepiece
- F. None of the Above

387. This definition means a component used in respirators to remove solid or liquid aerosols from the inspired air.

- A. Filter or Air-Purifying Element
- B. Filtering medium
- C. Escape Only Respirator
- D. End-Of-Service-Life Indicator
- E. Filtering Facepiece
- F. None of the Above

388. _____ means a particulate respirator with a filter as an integral part of the facepiece or with the entire facepiece composed of the filtering medium.

- A. Canister or Cartridge
- B. Air-Purifying Respirator
- C. Filtering Facepiece
- D. Disposable Respirators
- E. Demand Respirator
- F. None of the Above

389. This definition means a negative pressure particulate respirator with a filter as an integral part of the facepiece or with the entire facepiece composed of the filtering medium.

- A. Canister or Cartridge
- B. Air-Purifying Respirator
- C. Filtering Facepiece
- D. Disposable Respirators
- E. Demand Respirator
- F. None of the Above

390. _____ means a container with a filter, sorbent, or catalyst, or a combination of these items, which removes specific contaminants from the air passed through the container.

- A. Canister or Cartridge
- B. Air-Purifying Respirator
- C. Filtering Facepiece
- D. Disposable Respirators
- E. Demand Respirator
- F. None of the Above

391. This definition means a situation that requires the use of respirators due to the unplanned generation of a hazardous atmosphere (often of unknown composition) caused by an accident, mechanical failure, or other means and that requires evacuation of personnel or immediate entry for rescue or corrective action.

- A. Employee Exposure
- B. Atmosphere-Supplying Respirator
- C. Emergency Respirator Use Situation
- D. Emergency Situation
- E. Immediately Dangerous to Life or Health
- F. None of the Above

392. This definition means any occurrence such as, but not limited to, equipment failure, rupture of containers, or failure of control equipment that may or does result in an uncontrolled significant release of an airborne contaminant.

- A. Employee Exposure
- B. Atmosphere-Supplying Respirator
- C. Emergency Respirator Use Situation
- D. Emergency Situation
- E. Immediately Dangerous to Life or Health
- F. None of the Above

393. _____ means an exposure to a concentration of an airborne contaminant that would occur if the employee were not using respiratory protection.

- A. Employee Exposure
- B. Atmosphere-Supplying Respirator
- C. Emergency Respirator Use Situation
- D. Emergency Situation
- E. Immediately Dangerous to Life or Health
- F. None of the Above

394. This definition means a respirator that is discarded after the end of its recommended period of use, after excessive resistance or physical damage, or when odor breakthrough or other warning indicators render the respirator unsuitable for further use.

- A. Canister or Cartridge
- B. Air-Purifying Respirator
- C. Filtering Facepiece
- D. Disposable Respirators
- E. Demand Respirator
- F. None of the Above

395. This definition means a respirator that supplies the respirator user with breathing air from a source independent of the ambient atmosphere.

- A. Employee Exposure
- B. Atmosphere-Supplying Respirator
- C. Emergency Respirator Use Situation
- D. Emergency situation
- E. Immediately Dangerous to Life or Health
- F. None of the Above

396. The Employee is required to retain written information regarding medical evaluations, fit testing, and the respirator program.

- A. True
- B. False

397. Training will be provided prior to requiring the employee to use a respirator in the workplace.

- A. True
- B. False

The training shall ensure that each employee can demonstrate knowledge of at least the following:

398. Why the respirator is necessary and how to obtain an improper fit, usage, or maintenance can compromise the protective effect of the respirator.

- A. True
- B. False

399. Limitations and capabilities of OSHA.

- A. True
- B. False

Retraining shall be conducted annually and when:

400. Changes in the workplace or the type of respirator render previous training obsolete.

- A. True
- B. False